

COURSE PLAN					
ACADEMIC YEAR 2016-17					
PROGRAMME	:	<i>B.Sc. Chemistry</i>	LECTURE HOURS	:	36
SEMESTER	:	<i>1</i>	CREDITS	:	2
SUBJECT TITLE	:	Theoretical and Inorganic Chemistry-I	SUBJECT CODE	:	15U1CRCHE01
COURSE TEACHERS	:	<i>Dr. Franklin J (FJ), Dr. Ignatious Abraham(IGA)</i>			
Instructional Hours	:	<i>Two Instructional hours per week</i>			

Sessions	Topic/Module	Method of teaching *
1	Module 1-(Chemistry as a discipline of science): 3 h What is Science? - Scientific statements - Scientific methods - Observation - Posing a question - .	Conventional Teaching
2	Formulation of hypothesis - Experiment - Theory - Law - Revision of scientific theories and laws. Evolution of chemistry - Alchemy - Branches of chemistry	Conventional Teaching
3	Components of a research project - Introduction, review of literature, scope, materials and methods, results and discussion, conclusions and bibliography.	Conventional Teaching
4	Module II Basic Concepts in Chemistry (3h) Atomic mass and Molecular mass. Isotopes, isobars and isotones - Mole concept - Molar volume - Oxidation and reduction - Oxidation number and valency - Variable valency - Equivalent mass.	Conventional Teaching
5	Methods of expressing concentration: Weight percentage, molality, molarity, normality, formality, mole fraction, ppm and millimoles.	Conventional Teaching
6	Numerical Problems related to basic concepts.	Conventional Teaching
7	Module 3: Analytical Chemistry I (9 hrs) Quantitative Analysis. Primary standard-secondary standard, quantitative dilution -	Conventional Teaching
8-9	problems.	Conventional Teaching

10-11	Calibration of volumetric apparatus. Acid base titrations- titration curves – pH indicators.	Conventional Teaching
12	Redox titrations – Titration curve - Titrations involving KMnO_4 and $\text{K}_2\text{Cr}_2\text{O}_7$ - Redox indicators.	Conventional Teaching ASSIGNMENT I
13	Complexometric titrations – EDTA titrations - titration curves – metal ion indicators and characteristics.	Conventional Teaching
14	Errors in Chemical Analysis. Accuracy, precision, Types of error-absolute and relative error, methods of eliminating or minimizing errors.	Conventional Teaching ICT
15	Methods of expressing precision: mean, median, deviation, average deviation and coefficient of variation. Significant figures and its application.	Conventional Teaching
16	Module IV:Atomic Structure(9h) Introduction to atomic structure based on historical development – Rutherford's atom model and its limitations –	Conventional Teaching
17	Failure of classical physics – Black body radiation –	Conventional Teaching
18	Compton Effect - Planck's quantum hypothesis -	Conventional Teaching
19	Photoelectric effect	Conventional Teaching
20	Generalization of quantum theory - Atomic spectra of hydrogen and hydrogen like atoms – Ritz-combination principle–	Conventional Teaching
21	Bohr theory of atom – Calculation of Bohr radius, velocity and energy of an electron -	Conventional Teaching
22	Explanation of atomic spectra – Rydberg equation – Limitations of Bohr theory - Sommerfield modification -	Conventional Teaching
23	Louis de Broglie's matter waves – Wave-particle duality	Conventional Teaching ICT

24	Electron diffraction - Heisenberg's uncertainty principle	Conventional Teaching
25	<p>Module (V) Quantum Mechanical Model of Atom (12 hrs)</p> <p>Operator algebra – Linear and Hermitian operators –</p>	Conventional Teaching
26	Laplacian and Hamiltonian operators	Conventional Teaching ASSIGNMENT II
27	Eigen functions and Eigen values of an operator	Conventional Teaching
28	Postulates of quantum mechanics - Well behaved functions	Conventional Teaching
29	Time independent Schrödinger wave equation - Application to particle in a one dimensional box –	Conventional Teaching
30	–.Normalization of wave function - Particle in a three-dimensional box- Degeneracy.	Conventional Teaching ICT
31	Application of Schrödinger wave equation to hydrogen atom –	
32	Conversion of Cartesian coordinates to polar coordinates - The wave equation in spherical polar coordinates (derivation not required) -	Conventional Teaching
33	Radial and Angular functions (derivation not required) – Orbitals and concept of Quantum numbers (n, l, m).Radial functions - Radial distribution functions and their plots – Shapes of orbitals (s, p and d).	Conventional Teaching
34	Schrödinger equation for multi-electron atoms: Need for approximation methods.	Conventional Teaching
35	Electron spin – Spin quantum number - Pauli's Exclusion principle -	Conventional Teaching
36	Hund's rule of maximum multiplicity - Aufbau principle – Electronic configuration of atoms	Conventional Teaching

References

1. Jeffrey A. Lee, *The Scientific Endeavor: A Primer on Scientific Principles and Practice*, Pearson Education, 1999.
2. C.N.R. Rao, *Understanding Chemistry*, Universities Press India Ltd., Hyderabad, 1999.
3. Robert H. Hill and David Finster, *Laboratory Safety for Chemistry Students*, 1st Edition, Wiley, Hoboken, NJ, 2010.
4. M.C. Day and J. Selbin, *Theoretical Inorganic Chemistry*, East West Press, New Delhi, 2002.
5. B.R. Puri, L.R. Sharma and K.C. Kalia, *Principles of Inorganic Chemistry*, 31st Edition, Milestone Publishers and Distributors, New Delhi, 2013.
6. Satya Prakash, *Advanced Inorganic Chemistry, Volume 1*, 5th Edition, S. Chand and Sons, New Delhi, 2012.
7. J. Mendham, R.C. Denney, J. D. Barnes and M. Thomas, *Vogel's Text Book of Quantitative Chemical Analysis*, 6th Edition, Pearson Education, Noida, 2013.
8. A.K. Chandra, *Introductory Quantum Chemistry*, 4th Edition, Tata McGraw Hill Publishing Company, Noida, 1994.
9. R.K. Prasad, *Quantum Chemistry*, 4th Edition, New Age International(P) Ltd., New Delhi, 2012.
10. B.K. Sen, *Quantum Chemistry – Including Spectroscopy*, 3rd Edition, Kalyani publishers, New Delhi, 2010.

References

1. T.F Gieryn, *Cultural Boundaries of Science*, University of Chicago Press, Chicago, 1999.
2. H. Collins and T. Pinch, *The Golem: What Everyone Should Know about Science*, Cambridge University Press, Cambridge, 1993.
3. C.R. Kothari, *Research Methodology: Methods and Techniques*, 2nd Revised Edition, New Age International Publishers, New Delhi, 2004.
4. *Guidance in a Nutshell - Compilation of Safety Data Sheets*, European Chemicals Agency, Finland, Version 1.0, December 2013.
5. J. D. Lee, *Concise Inorganic Chemistry*, 5thedn., Blackwell Science, London (Chapter 1)
1. D. F. Shriver and P. W. Atkins, *Inorganic Chemistry*, 3rdedn., Oxford University Press(Chapter 1)
2. B. Douglas, D. Mc Daniel, J. Alexander, *Concepts and models in Inorganic Chemistry* (Chapter 1)

3. D.A. Skoog, D.M. West, F.J. Holler and S.R. Crouch, *Fundamentals of Analytical Chemistry*, 8th Edition, Brooks/Cole, Thomson Learning, Inc., USA, 2004
4. D.A. McQuarrie, *Quantum Chemistry*, 2nd Edition, University Science Books, California, 2008.
5. M.C. Day and J. Selbin, *Theoretical Inorganic Chemistry*, East West Press, New Delhi, 2002.
11. P.W. Atkins and R.S. Friedman, *Molecular Quantum Mechanics*, 3rd Edition, Oxford University Press, New York, 1997.
12. I.N. Levine, *Quantum Chemistry*, 6th Edition, Pearson Education Inc., New Delhi, 2009.
13. Jack Simons, *An Introduction to Theoretical Chemistry*, 2nd Edition, Cambridge University Press, Cambridge, 2005.

COURSE PLAN				
ACADEMIC YEAR 2016-17				
PROGRAMME	:	<i>B.Sc. Chemistry</i>	LECTURE HOURS	: 36
SEMESTER	:	2	CREDITS	: 2
SUBJECT TITLE	:	<i>Theoretical and Inorganic Chemistry-II</i>	SUBJECT CODE	: 15U2CRCHE02
COURSE TEACHERS	:	<i>Dr. Franklin J(FJ), Dr. Ignatious Abraham(IGA)</i>		
COURSE OUTCOMES (COs)	:	<p>CO1 - Ability to understand the basics of periodicity in the properties of the elements, chemical bonding, nuclear chemistry and different analytical techniques</p> <p>CO2 - Ability to apply valence bond and molecular orbital theories to explain the bonding characteristics of different chemical systems.</p> <p>CO3 - Ability to interpret the properties such as dipole moment, bond length, magnetic behaviour and bond energy of molecular systems in the light of VB or MO theory.</p> <p>CO4 - Ability to explore and reflect about the wide range of possibilities and applications of nuclear reactions and radio activity.</p> <p>CO5 - Ability to apply gravimetric analysis and different separation/purification techniques effectively in laboratory scale.</p>		
Instructional Hours	:	<p><i>Monday : Period 1 (9:30 to 10:30 am)</i></p> <p><i>Tuesday: Period 5 (2:30 to 3:30 pm)</i></p>		
IGA	No. of	Session Topic and Discussion Theme	Value additions	COs

	Session			
UNIT 2 : Chemical Bonding -1 (9 hours)	1	Introduction – Type of bonds – Octet rule and its limitations.		CO1
	2	<i>Ionic Bond</i> : Factors favoring the formation of ionic bonds - Lattice energy of ionic compounds - Born-Lande equation (derivation not expected) Solvation enthalpy and solubility of ionic compounds		CO1
	3	Born-Haber cycle and its applications –Properties of ionic compounds - Polarisation of ions – Fajan'srules and its applications.	Assignment No: 1	CO1
	4	<i>Covalent Bond</i> : Lewis theory. Valence Bond Theory. Co ordinate bond		CO1
	5	Hybridization: Definition and characteristics VSEPR theory: Postulates		CO1
	6	Applications – Shapes of molecules- sp (BeCl ₂ , C ₂ H ₂), sp ² (BF ₃ , C ₂ H ₄), sp ³ (CH ₄ , CCl ₄ , NH ₃ , H ₂ O, NH ₄ ⁺ , H ₃ O ⁺ and SO ₄ ²⁻)		CO1, CO2
	7	sp ³ d (PCl ₅), sp ³ d ² (SF ₆) and sp ³ d ³ (IF ₇) and SF ₄ , ClF ₃ , XeF ₂ , IF ₅ , XeF ₄ , IF ₇ and XeF ₆	MOODLE- Assignment No:2	CO1, CO2
	8	Limitations of VBT. Properties of covalent compounds. Polarity of covalent bond – Percentage of ionic character –Dipole moment and molecular structure.		CO1, CO2, CO3
	9	FIRST INTERNAL EXAMINATION		
Text Books	<ul style="list-style-type: none"> ❖ B.R. Puri, L.R. Sharma and K.C. Kalia, <i>Principles of Inorganic Chemistry</i>, 31st Edition, Milestone Publishers and Distributors, New Delhi, 2013. ❖ Satya Prakash, <i>Advanced Inorganic Chemistry, Volume 1</i>, 5th Edition, S. Chand and Sons, NewDelhi, 2012. ❖ Manas Chanda, <i>Atomic Structure and Chemical Bonding</i>, 4th Edition, Tata McGraw Hill 			
periodic properties	10	Modern periodic law – Long form periodic table. Periodicity in properties: Atomic and ionic radii	Assignment No.3	CO1
	11	Ionization enthalpy - Electron affinity (electron gain enthalpy) – Electronegativity. Electronegativity scales: Pauling and Mullikan scales		CO1

(4 hours)	No. of Session	Session Topic and Discussion Theme	Value additions	COs
	12	Effective nuclear charge – Slater rule and its applications		CO1
	13	Revision-Periodicity in properties and its consequences	Group Discussion	CO1, CO3
SECOND INTERNAL EXAMINATION				
<i>Text Books</i>	<ul style="list-style-type: none"> ❖ B.R. Puri, L.R. Sharma and K.C. Kalia, <i>Principles of Inorganic Chemistry</i>, 31st Edition, Milestone Publishers and Distributors, New Delhi, 2013. ❖ Satya Prakash, <i>Advanced Inorganic Chemistry, Volume 1</i>, 5th Edition, S. Chand and Sons, NewDelhi, 2012. ❖ Manas Chanda, <i>Atomic Structure and Chemical Bonding</i>, 4th Edition, Tata McGraw Hill 			
UNIT 5 : Analytical Chemistry II (5 hours)	14	Gravimetric analysis: Systematic steps in gravimetric analysis. Illustrations using iron and barium estimation.		CO1
	15	Separation and purification techniques – Filtration, Crystallization and precipitation – Fractional distillation, Solvent extraction.		CO1, CO5
	16	Concept of solubility product as applied in group separation of cations – problems.	Demonstration	CO1, CO5
	17	Chromatography - Classification of methods elementary study of adsorption, paper, thin layer, column, ion exchange chromatography	PowerPoint presentation	CO1, CO5
	18	Gas chromatographic methods. HPLC	PowerPoint presentation	CO1, CO5
<i>Text Books</i>	<ul style="list-style-type: none"> ❖ Vogel’s Textbook of Quantitative Chemical Analysis 6th edn, Pearsons Education Ltd ❖ R. D. Day, A. L. Underwood, Quantitative analysis, 6th Edn., Prentice Hall of India Pvt. Ltd ❖ Satya Prakash, <i>Advanced Inorganic Chemistry, Volume 1</i>, 5th Edition, S. Chand and Sons, NewDelhi, 2012. 			

FJ					
<i>Unit IV</i>	No. of Sessions	Session Topic and Discussion Theme	Value additions	COs	
Nuclear Chemistry	1	Introduction to nuclear chemistry Structure of nucleus ➤ Nuclear particles, nuclear forces, nuclear size, nuclear density		CO1	
	2	Stability of nucleus ➤ binding energy ➤ magic numbers ➤ packing fraction ➤ n/p ratio. Nuclear Models	Individual Assignment: Nuclear models – liquid drop model and shell model.	CO1, CO4	
	3	Natural Radioactivity ➤ modes of decay, decay constant ➤ half-life period, average life	Problems: Based on radioactive decay	CO1	
	4	Radioactive Equilibrium Geiger-Nuttal rule, units of radioactivity, radiation dosage		CO1	
	5	Nuclear Reactions ➤ induced by charged projectiles, neutrons and γ rays		CO1, CO4	
	7	fission reactions fusion reactions		CO1, CO4	
	1st Internal Examination				
	8	Preparation of transuranic elements		CO1	
	9	Chain Reactions, Stellar energy	Power Point Presentation	CO1	

<i>Text Books</i>	<ul style="list-style-type: none"> ❖ H. J. Arnika, Essentials of Nuclear Chemistry, New Age ❖ R. Gopalan, Elements of Nuclear Chemistry, Vikas Publ. House. ❖ B.R. Puri, L.R. Sharma and K.C. Kalia, <i>Principles of Inorganic Chemistry</i>, 31st Edition, Milestone Publishers and Distributors, New Delhi, 2013. ❖ Satya Prakash, <i>Advanced Inorganic Chemistry, Volume 1</i>, 5th Edition, S. Chand and Sons, New Delhi, 2012. ❖ Manas Chanda, <i>Atomic Structure and Chemical Bonding</i>, 4th Edition, Tata McGraw Hill Publishing Company, Noida, 2007. 			
Unit III	No. of Sessions	Session Topic and Discussion Theme	Value additions	COs
<i>Chemical Bonding – II</i>	10	MO Theory <ul style="list-style-type: none"> ➤ Linear combination of atomic orbitals ➤ Formation of molecular orbitals ➤ Bonding and antibonding molecular orbitals ➤ Stability of molecules based on bond order ➤ Relation between bond order and bond length 		CO1, CO2, CO3
	11	MO diagram of homo nuclear system <ul style="list-style-type: none"> ➤ H₂, He₂, Li₂, Be₂, B₂, C₂, N₂, O₂, F₂ ➤ Magnetic behaviour of these homo nuclear systems 	Power Point Presentation	CO1, CO2, CO3
	12	MO diagram of heteronuclear system <ul style="list-style-type: none"> ➤ CO and NO ➤ Magnetic behaviour of these homo nuclear systems ➤ Comparison of bond length, magnetic behaviour and bond energy of O₂, O₂⁺, O₂²⁺, O₂⁻ and O₂²⁻ 	Power Point Presentation	CO1, CO2, CO3
	13	Resonance structures of <ul style="list-style-type: none"> ➤ borate, carbonate and nitrate ions ➤ Comparison of bond energy. 		CO1
	14	Comparison of VB and MO theories.	Group Discussion	CO1

2nd Internal Examination			
	15	Metallic Bond ➤ Free electron theory ➤ valence bond theory	CO1
	16	Band theory Explanation of metallic properties based on these theories.	CO1
	17	Intermolecular Forces ➤ Induction forces and dispersion forces	Individual Assignment: Various intermolecular forces
	18	Hydrogen bond Intra and inter molecular hydrogen bonds , Effect on physical properties.	CO1
Text Books	<ul style="list-style-type: none"> ❖ B.R. Puri, L.R. Sharma and K.C. Kalia, <i>Principles of Inorganic Chemistry</i>, 31st Edition, Milestone Publishers and Distributors, New Delhi, 2013. ❖ Satya Prakash, <i>Advanced Inorganic Chemistry, Volume 1</i>, 5th Edition, S. Chand and Sons, NewDelhi, 2012. ❖ Manas Chanda, <i>Atomic Structure and Chemical Bonding</i>, 4th Edition, Tata McGraw Hill 		

DEPARTMENT OF CHEMISTRY, SACRED HEART COLLEGE (AUTONOMOUS), THEVARA

COURSE PLAN : ACADEMIC YEAR 2016 - 2017

PROGRAMME	: <i>B.Sc. Chemistry</i>	SEMESTER	: 3
LECTURE HOURS	: 54	CREDITS	: 3
SUBJECT TITLE	: <i>Organic Chemistry - I</i>	SUBJECT CODE	: 15U3CRCHE03

COURSE TEACHERS	: Dr. Joseph John (JJ), Dr. V. S. Sebastian (VSS) & Dr. Grace Thomas (GT)
COURSE OUTCOMES (COs)	: <ul style="list-style-type: none"> CO1 - <i>Understand the classification and nomenclature of organic compounds.</i> CO2 - <i>Describe aromaticity and stereochemistry of organic compounds.</i> CO3 - <i>Understand the fundamentals of organic reaction mechanisms.</i> CO4 - <i>Describe various emerging areas of organic chemistry and its applications.</i>
Instructional Hours	: <ul style="list-style-type: none"> <i>Tuesday : Period 1 (9:30 to 10:30 am) - JJ</i> <i>Wednesday : Period 1 (9:30 to 10:30 am) - FJ</i> <i>Thursday : Period 5 (2:30 to 3:30 pm) - VSS</i>

GRACE THOMAS

Unit I : CLASSIFICATION AND NOMENCLATURE OF ORGANIC COMPOUNDS

Session s	Session Topic and Discussion Theme	Value additions	COs
1	Introduction to classification of organic compounds.		CO1
2	Rules of IUPAC system of nomenclature		CO1
3	Alkanes, alkenes, alkynes, cycloalkanes, bicycloalkanes, alkyl halides, alcohols and phenols.	Individual Assignment: <i>Nomenclature of Compounds</i>	CO1
4	Aldehydes, ketones, carboxylic acids and its derivatives, amines, nitro compounds. (Both aliphatic and aromatic)		CO1

Unit IV: AROMATICITY

5	Concept of resonance: <ul style="list-style-type: none"> • Resonance energy in benzene. • Heat of hydrogenation and heat of combustion of Benzene • C-C bond lengths and orbital picture of Benzene 	Power Point Presentation Structure and resonance of Benzene	CO2
6	Structure of naphthalene and anthracene (<i>Molecular Orbital picture and resonance</i>)		CO2
7	Concept of aromaticity – aromaticity (definition), <i>Huckel's rule:</i> Application to Benzenoids : Benzene, Naphthalene, Anthracene, Phenanthrene.		CO2
8	Non-Benzenoid compounds – cyclopropenyl cation, cyclopentadienyl anion and tropylium cation.		CO2
9	General mechanism of electrophilic substitution-mechanism of halogenation, nitration,		CO2

1 st Internal Examination			
10	Mechanism of Friedal Craft's alkylation and acylation, sulphonation.		CO2
11	Orientation of aromatic substitution – <i>ortho</i> , <i>para</i> and <i>meta</i> directing groups.		CO2
12	Ring activating and deactivating groups.	Power Point Presentation	CO2
13	Electronic interpretation of various groups like - NO ₂ and –OH		CO2
14	Orientation (i). Amino, methoxy and methyl groups (ii). Carboxy, nitro, nitrile, carbonyl and sulfonic acid groups. (iii). Halogens.	Power Point Presentation: <i>Illustration of orientation effects</i>	CO2
15	Reactivity of naphthalene towards electrophilic substitution. Nitration and sulphonation		CO2

2nd Internal Examination

16	Aromatic Nucleophilic substitutions - bimolecular displacement mechanism		CO2
17	Elimination –Addition mechanism : <i>Benzyne intermediate</i>	Power Point Presentation: <i>Benzyne intermediates</i>	CO2
18	Reactivity and orientation in Aromatic Nucleophilic substitutions.		CO2

References:

1. R. T. Morrison and R.N Boyd, 'Organic Chemistry', 6th Edition - Prentice Hall of India.
2. L. Finar, 'Organic Chemistry' - Vol.- 6th Edition I, Pearson Education (chapters-20,21)
3. M. K. Jain and S.C. Sharma 'Modern Organic Chemistry', 3rd Edn, Vishal Publishing Co.
4. K. S. Tewari and N. K. Vishnoi 'Organic Chemistry', 3rd Edition, Vikas Publishing House,
5. Peter Sykes, A Guide book to Mechanism in Organic Chemistry ;, 6th Edition, Pearson Education.

JOSEPH JOHN

Unit II : ORGANIC REACTION MECHANISMS

Session s	Session Topic and Discussion Theme	Value additions	COs
1	Introduction to Organic Reaction Mechanisms		CO3
2	Drawing electron movements with arrows: <ul style="list-style-type: none"> • Curved arrow notation. • Half headed and double headed arrows. 	Individual Assignment: <i>Reactions with different kinds of arrow notations</i>	CO3
3	<i>Types of reagents:</i> Electrophiles and Nucleophiles <i>Types and sub-types organic reactions:</i> Substitution, Addition reactions, Elimination and Rearrangement.	Classification of 20 named reactions	CO3
4	Reactive intermediates with examples – carbocations, carbanions	Power Point Presentation: <i>Structure of intermediates</i>	CO3
5	Reactive intermediates with examples - carbenes, nitrenes and free radicals.		CO3
6	Electron displacement effects - Inductive, electrometric, mesomeric, resonance		CO3
7	Hyperconjugation and steric effects- steric inhibition of resonance.		CO3
8	<ul style="list-style-type: none"> • Aliphatic nucleophilic substitutions, 		CO3

	<ul style="list-style-type: none"> mechanism of S_N1 and S_N2 reactions. 		
9	<ul style="list-style-type: none"> Effects of structure, substrate, solvent, nucleophile and leaving groups Stereochemistry- Walden inversion 		CO3
1st Internal Examination			
10	<i>Elimination Reactions</i> :-Hoffmann and Saytzeff rules		CO3
11	<i>Cis</i> - and <i>trans</i> - eliminations		
12	<ul style="list-style-type: none"> Mechanisms of E1 and E2 reactions Elimination <i>versus</i> substitution. 	Power Point Presentation	CO3
13	Addition reactions: <ul style="list-style-type: none"> Mechanisms of addition of Bromine Inductomeric effect 		CO3
14	Mechanisms of addition of hydrogen halides to double bonds.		CO3
15	Markonikoff's rule and peroxide effect.		CO3
2nd Internal Examination			
16	Polymerisation reactions: Types of polymerization - free radical, cationic and anionic – polymerisations –including mechanism.		CO4
17	Pericyclic Reactions: Classification- electrocyclic, sigmatropic, cycloaddition reactions-Examples	Power Point Presentation: <i>Illustration of different reactions</i>	CO4
18	Diels- Alder reaction- Stereochemical aspects- Effect of substituents.		CO4

References:

- Peter Sykes, A Guide book to Mechanism in Organic Chemistry: 6th Edition, Pearson Education.
- P. S. Kalsi 'Organic Reactions and their Mechanisms' New Age International Publishers.
- K.S. Tewari and N.K. Vishnoi 'Organic Chemistry', 3rd Edition, Vikas Publishing House.
- M. K. Jain and S.C. Sharma 'Modern Organic Chemistry', 3rd Edition, Vishal Publishing Company Co.
- R. T. Morrison and R. N. Boyd, 'Organic Chemistry', 6th Edition - Prentice Hall of India,
- I. L. Finar, Organic Chemistry, 6th Edition. Vol.- I, Pearson

V S SEBASTIAN**Unit III : STEREOCHEMISTRY OF ORGANIC COMPOUNDS**

Session s	Session Topic and Discussion Theme	Value additions	COs
1	Stereoisomerism - definition - classification - optical and geometrical isomerism		CO2
2	Projection formulae - Fischer, flying wedge, Sawhorse and Newman projection formulae - notation of optical isomers -D-L notation-	Power point presentation	CO2
3	Cahn-Ingold-Prelog rules - R-S notations for optical isomers with one and two asymmetric carbon atoms - erythro and threo representations.		CO2
4	Optical isomerism - optical activity - optical and specific rotations - conditions for optical activity	Power Point Presentation: <i>Structure of intermediates</i>	CO2
5	Asymmetric centre: Chirality - achiral molecules - meaning of (+) and (-)		CO2
6	Elements of symmetry -. Prochirality		CO2

7	Racemization - methods of racemization (by substitution and tautomerism)		CO2
8	Resolution - methods of resolution -mechanical, seeding, biochemical and conversion to diastereoisomers		CO2
9	Asymmetric synthesis (partial and absolute synthesis).		CO2
1st Internal Examination			
10	Optical activity in compounds does not containing asymmetric carbon atoms-Biphenyls and allenes.		CO2
11	Geometrical isomerism - <i>cis-trans syn-anti</i> and E-Z notations - geometrical isomerism in maleic and fumaric acids and unsymmetrical ketoximes		CO2
12	Methods of distinguishing geometrical isomers using melting point, dipole moment, dehydration and cyclisation.	Power Point Presentation	CO2
13	Conformational analysis - introduction of terms - conformers, configuration, dihedral angle, torsional strain	Power Point Presentation	CO2
14	Conformational analysis of ethane and <i>n</i> -butane using energy profile diagrams	Power Point Presentation	CO2
15	Conformers of cyclohexane (chair, boat and skew boat forms) - axial and equatorial- bonds-ring flipping showing axial equatorial interconversion,	Power Point Presentation	CO2
2nd Internal Examination			
16	Conformation of methyl cyclohexane.		
Unit V : SUPRAMOLECULAR CHEMISTRY			
17	Introduction to Supramolecular chemistry	Power Point Presentation:	CO4
18	Structure of supramolecular compounds		CO4
References:			
<ol style="list-style-type: none"> 1. P. Sykes, A Guide book to Mechanism in Organic Chemistry, 6th Edition, Orient Longman. 2. P.S. Kalsi' 'Organic Reactions and their Mechanisms'' New Age International Publishers. 3. M. K. Jain and S. C. Sharma 'Modern Organic Chemistry', 3rd Edition, Vishal Publishing Co. 4. P. Y. Bruice, 'Organic Chemistry' - 3rd Edn. Pearson Education. 			

DEPARTMENT OF CHEMISTRY, SACRED HEART COLLEGE (AUTONOMOUS), THEVARA	
COURSE PLAN : ACADEMIC YEAR 2016 - 2017	
PROGRAMME	: <i>B.Sc. Chemistry</i> SEMESTER : 4
LECTURE HOURS	: 54 CREDITS : 3
SUBJECT TITLE	: <i>Organic Chemistry - I</i> SUBJECT CODE : 15U4CRCHE04
COURSE TEACHERS	: Dr. Joseph John (JJ), Dr. V. S. Sebastian (VSS) & Dr. Grace Thomas (GT)
COURSE OUTCOMES (COs)	: CO1 - <i>Understand the chemistry of some selected functional groups with a view to develop proper aptitude towards the study of organic compounds and their reactions.</i> CO2 - <i>Describe the chemistry of alcohols, phenols, carboxylic acids, derivatives of Carboxylic acids, Sulphonic acids, carbonyl compounds, poly nuclear hydrocarbons, active methylene compounds and Grignard reagents</i> CO3 - <i>Understand and explain organic reaction mechanisms.</i>
Instructional Hours	: <i>Tuesday : Period 1 (9:30 to 10:30 am) - JJ</i> <i>Wednesday : Period 1 (9:30 to 10:30 am) - VSS</i> <i>Thursday : Period 5 (2:30 to 3:30 pm) - GT</i>

GRACE THOMAS			
Unit I : HYDROXY COMPOUNDS			
Session s	Session Topic and Discussion Theme	Value additions	COs
1	Introduction to Alcohols.		CO1
2	Monohydric alcohols: Classification Physical properties–hydrogen bonding		CO1
3	Distinction between primary, secondary and tertiary alcohols. Ascent and decent in alcohol series	Individual Assignment: <i>Nomenclature of Alcohols</i>	CO1
4	Dihydric alcohols		CO1
5	Oxidative cleavage – Lead tetra acetate, Periodic acid Pinacol - Pinacolone rearrangement: mechanism	Power Point Presentation	CO3
6	Phenols – Acidity of phenols- effects of substituents – comparison of acidity with alcohols		CO2
7	Preparation and uses of nitrophenols, picric acid, catechol, resorcinol and quinol		CO2
8	Mechanisms of Reimer –Tiemann reaction, Lederer- Mannase reaction, Fries rearrangement		CO3
Unit VI : GRIGNARD AND RELATED COMPOUNDS			
9	Grignard reagents-formation, structure and synthetic applications		CO2
1 st Internal Examination			
10	Alkyl lithium, Reformatsky reaction		CO2

Unit VII : COMPOUNDS CONTAINING ACTIVE METHYLENE GROUPS			
11	Synthetic uses of malonic ester.	<i>Assignment :</i> Synthetic Applications of active methylene compounds	CO2
12	Synthetic uses of acetoacetic ester.		CO2
13	Synthetic uses of cyanoacetic ester.		CO2
14	Keto-enol tautomerism of ethyl acetoacetate.		CO2
15	Alkylation of carbonyl compounds <i>via</i> enamines		CO2
2nd Internal Examination			
Unit VIII : POLY NUCLEAR HYDROCARBONS AND THEIR DERIVATIVES			
16	Classification – reactions and structure of naphthalene	Power Point Presentation	CO2
17	Reactions and structure of anthracene and phenanthrene		CO2
18	Elementary idea of naphthyl amines, naphthols, naphthaquinone and anthraquinone.		CO2
References:			
<ol style="list-style-type: none"> 1. R. T. Morrison and R.N Boyd, 'Organic Chemistry', 6th Edition - Prentice Hall of India. 2. L. Finar, 'Organic Chemistry' - Vol.- 6th Edition I, Pearson Education (chapters-20,21) 3. M. K. Jain and S.C. Sharma 'Modern Organic Chemistry', 3rd Edn, Vishal Publishing Co. 4. K. S. Tewari and N. K. Vishnoi 'Organic Chemistry', 3rd Edition, Vikas Publishing House, 5. Peter Sykes, A Guide book to Mechanism in Organic Chemistry ;, 6th Edition, Pearson Education. 			
V S SEBASTIAN			
Unit II : ETHERS AND EPOXIDES			
Session s	Session Topic and Discussion Theme	Value additions	COs
1	Synthesis and Reactions of Epoxides		CO2
2	Cleavage of ether linkages by HI- Ziesels method of estimation of alkoxy groups	Power point presentation	CO2
3	Claisen rearrangement –mechanism.		CO2
Unit III : ALDEHYDES AND KETONES			
4	Structure and reactivity of the carbonyl group - acidity of alpha hydrogen. Comparative studies of -aldehydes and ketones -	Power Point Presentation:	CO2
5	aliphatic and aromatic aldehydes - formaldehyde and acetaldehyde-		CO2
6	Mechanism of nucleophilic additions to carbonyl groups : Claisen and Claisen-Schmidt condensations.		CO2
7	Mechanism of nucleophilic additions to carbonyl groups : Benzoin condensations and Aldol condensations.	Assignment: Mechanism of Named reactions	CO2
8	Mechanism of nucleophilic additions to carbonyl groups : Perkin and Knoevenagel condensations.		CO2
9	Condensation with ammonia and its derivatives.		CO2
1st Internal Examination			
10	Wittig reaction and Mannich reaction, Addition of Grignard reagents.		CO3
11	Oxidation and reduction of aldehydes and ketones: Baeyer-Villiger oxidation		CO3

12	Oxidation and reduction of aldehydes and ketones: Cannizzaro's reaction		CO3
13	Oxidation and reduction of aldehydes and ketones: Meerwein-Ponndorf -Verley, and Clemmensen, reductions		CO3
14	Oxidation and reduction of aldehydes and ketones: Wolff-Kishner, LiAlH ₄ and NaBH ₄ reductions.		CO3
15	Use of acetal as protecting group.	Power Point Presentation	CO2
2nd Internal Examination			
Unit V : CARBONIC ACID DERIVATIVES			
16	Preparation, reactions, structure and uses: Urea and thiourea		CO2
17	Preparation, reactions, structure and uses: semicarbazide		CO2
18	Preparation and basicity of guanidine.		CO2
References:			
<ol style="list-style-type: none"> 1. P. Sykes, A Guide book to Mechanism in Organic Chemistry, 6th Edition, Orient Longman. 2. P.S. Kalsi' 'Organic Reactions and their Mechanisms'' New Age International Publishers. 3. M. K. Jain and S. C. Sharma 'Modern Organic Chemistry', 3rd Edition, Vishal Publishing Co. 4. P. Y. Bruice, 'Organic Chemistry' - 3rd Edn. Pearson Education. 			
JOSEPH JOHN			
Unit IV : CARBOXYLIC AND SULPHONIC ACIDS			
Session s	Session Topic and Discussion Theme	Value additions	COs
1	Structure of carboxylate ion- effects of substituents on acid strength of aliphatic and aromatic carboxylic acids	Power Point Presentation	CO2
2	Ascent and descent in fatty acid series		CO2
3	Hell-Volhard-Zelinsky reaction -Mechanism of decarboxylation		CO3
4	Preparation of functional derivatives of carboxylic acids : acid chlorides		CO2
5	Preparation of functional derivatives of carboxylic acids : esters and anhydrides		CO2
6	Preparation of functional derivatives of carboxylic acids: amides		CO2
7	Methods of formation and chemical reactions : anthranilic acid, cinnamic acid		CO2
8	Methods of formation and chemical reactions : acrylic acid		CO2
9	Methods of formation and chemical reactions : oxalic acid	Assignment	CO2
1st Internal Examination			
10	Methods of formation and chemical reactions : malonic acid		CO2
11	Methods of formation and chemical reactions : citric acid		CO2

12	Methods of formation and chemical reactions: adipic acid		CO2
13	Methods of formation and chemical reactions: maleic acid,		CO2
14	Methods of formation and chemical reactions: fumaric acid		CO2
15	Methods of formation and chemical reactions: coumarin.		CO2
2nd Internal Examination			
16	Preparation, reactions and uses: benzene sulphonic acid		CO2
17	Preparation, reactions and uses: benzene sulphonyl chloride		CO2
18	Preparation, reactions and uses: <i>ortho</i> and <i>para</i> toluene sulphonyl chlorides		CO2
References:			
<ol style="list-style-type: none"> 1. Peter Sykes, A Guide book to Mechanism in Organic Chemistry: 6th Edition, Pearson Education. 2. P. S. Kalsi' 'Organic Reactions and their Mechanisms'' New Age International Publishers. 3. K.S. Tewari and N.K. Vishnoi 'Organic Chemistry', 3rd Edition, Vikas Publishing House. 4. M. K. Jain and S.C. Sharma 'Modern Organic Chemistry', 3rd Edition, Vishal Publishing Company Co. 5. R. T. Morrison and R. N. Boyd, 'Organic Chemistry', 6th Edition - Prentice Hall of India, 6. I. L. Finar, Organic Chemistry, 6th Edition. Vol.- I, Pearson 			

COURSE PLAN

ACADEMIC YEAR 2016-17

PROGRAMME	:	<i>B.Sc. Chemistry</i>	LECTURE HOURS	:	<i>54</i>
SEMESTER	:	<i>5</i>	CREDITS	:	<i>3</i>
SUBJECT TITLE	:	<i>Inorganic Chemistry - I</i>	SUBJECT CODE	:	<i>15U5CRCHE05</i>
COURSE TEACHERS	:	<i>Dr. Joseph John (JJ), Mr. Midhun Dominic C D (MD), Ms. June Cyriac (JUC)</i>			
Objectives	:	<ul style="list-style-type: none"> ➤ To understand the general characteristics of the d and f block elements ➤ To study the physical and chemical properties of d and f block elements ➤ To study the various theories of coordination compounds ➤ To study isomerism exhibited by metal complexes ➤ To understand the classification, properties and applications of organometallic compounds. ➤ To study the concepts of acids and bases. ➤ To understand the importance and various functions of metals in biological systems. 			
Instructional Hours	:	<i>3 hours per week</i>			

JJ	No. of Session	Session Topic and Discussion Theme	Value additions	Remarks
UNIT 2: Co-	1	Ligands, classification based on the number of donor atoms, chelating ligands, bridging ligands		
	2	Isomerism in complexes-Structural isomerism only, IUPAC nomenclature		

<i>ordination Chemistry (18 hours)</i>	3	Coordination number and possible geometries of complexes from C. N 3 to 12	Assignment No: 1	
	4	Stereo isomerism- geometrical and optical isomerism of complexes with 4 and 6 coordination numbers.		
	5	Stability of complexes - stepwise stability constant and overall stability constant, factors affecting the stability of metal complexes. EAN, Chelates and chelate effect.	Group Discussion	
	6	Magnetic behavior of complexes- Diamagnetic and paramagnetic complexes, explanation, effective magnetic moment, spins only magnetic moments, calculation of spin only magnetic moment. Quenching of magnetic moment.		
	7	Theories of bonding in coordination compounds– Werner’s theory of coordination, primary and secondary valences of metal ions.	MOODLE- Assignment No:2	
	8	Valence bond theory- of octahedral, tetrahedral and square planar complexes, high spin and low spin complexes- inner and outer orbital complexes, explanation of magnetic properties		
	9	FIRST INTERNAL EXAMINATION		
	<i>Text Books</i>	<ul style="list-style-type: none"> ❖ J. D. Lee, Concise Inorganic Chemistry 5th edn., Wiley India Pvt. Ltd.2008. ❖ R. Puri, L. R. Sharma, K. C. Kalia, Principles of Inorganic Chemistry,31st Milestone Publishers, New Delhi 2010 ❖ G. L. Meissler, D.A Tarr, Inorganic Chemistry,3rd Edn. Pearson Education, 2004. ❖ J. E. Huheey, E. A. Keiter, R. L. Keiter, O K Medhi, Inorganic Chemistry, Pearson 2006 ❖ F. A. Cotton and G. Wilkinson, Advanced Inorganic Chemistry 6th edn., John Wiley, New York 1991. 		
<i>UNIT 2 : Co-</i>	10	Limitations of valence bond theory.	Assignment No.3	
	11	Crystal field theory- splitting of d-orbitals in octahedral, tetrahedral complexes, strong and weak field ligands, pairing energy		

<i>ordinati on Chemistry (18 hours)</i>	12	Explanation of colour and magnetic properties of complexes, limitation of CF theory.		
	13	Jahn-Teller distortion and splitting of d orbitals in tetragonal and square planar fields,		
	14	Jahn-Teller distortion in Cu (II) complexes. MO theory- evidence for metal ligand covalency		
		SECOND INTERNAL EXAMINATION		
<i>Text Books</i>	<ul style="list-style-type: none"> ❖ J. D. Lee, Concise Inorganic Chemistry 5th edn., Wiley India Pvt. Ltd.2008. ❖ R. Puri, L. R. Sharma, K. C. Kalia, Principles of Inorganic Chemistry,31st Milestone Publishers, New Delhi 2010 ❖ G. L. Meissler, D.A Tarr, Inorganic Chemistry,3rd Edn. Pearson Education, 2004. ❖ J. E. Huheey, E. A. Keiter, R. L. Keiter, O K Medhi, Inorganic Chemistry, Pearson 2006 ❖ F. A. Cotton and G. Wilkinson, Advanced Inorganic Chemistry 6th edn., John Wiley, New York 1991. 			
<i>UNIT 2 : Co-ordination Chemistry (18 hours)</i>	15	MOE diagram of complexes of octahedral symmetry (sigma bonding only).		
	16	Explanation of Δ in the Oh and Td complexes using MOE diagram.		
	17	Substitution reactions in metal complexes- Labile and inert complexes, ligand substitution reactions in octahedral complexes	Demonstration	
	18	S _N 1 and S _N 2 substitution reactions. Substitution reactions of square planar complexes – Trans effect and applications of trans effect.	PowerPoint presentation	
<i>Text Books</i>	<ul style="list-style-type: none"> ❖ J. D. Lee, Concise Inorganic Chemistry 5th edn., Wiley India Pvt. Ltd.2008. ❖ R. Puri, L. R. Sharma, K. C. Kalia, Principles of Inorganic Chemistry,31st Milestone Publishers, New Delhi 2010 ❖ G. L. Meissler, D.A Tarr, Inorganic Chemistry,3rd Edn. Pearson Education, 2004. ❖ J. E. Huheey, E. A. Keiter, R. L. Keiter, O K Medhi, Inorganic Chemistry, Pearson 2006 ❖ F. A. Cotton and G. Wilkinson, Advanced Inorganic Chemistry 6th edn., John Wiley, New York 1991. 			

MD	No. of Session	Session Topic and Discussion Theme	Value additions	Remarks
UNIT 4 : Acids and Bases (8 hours)	1	Definition- Bronsted-Lowry, Lux-Flood, Solvent system, Usanovich and Lewis definitions. Self-study.		
	2	Strength of lewis acids and bases: Factors affecting strength of acids and bases: Solvent effect		
	3	Factors affecting strength of acids and bases: Leveling and differentiating solvents.	Assignment No: 1	
	4	Effect of substituent, steric factor, charge on the species, Electronegativity		
	5	Hydration, oxidation number of the central atom, resonance effect	Group Discussion	
	6	Hard and Soft acids and bases. HSAB Theory, basis of HSAB theory		
	7	Symbiosis, Applications of HSAB Concept	MOODLE- Assignment No:2	
	8	Stability of complexes, mode of coordination, predicting feasibility of reactions		

FIRST INTERNAL EXAMINATION			
Text Books	<ul style="list-style-type: none"> ❖ B. R. Puri, L. R. Sharma, K. C. Kalia, Principles of Inorganic Chemistry, 31st Edn. Milestone Publishers, New Delhi 2010. ❖ G. L. Meissler, D. A Tarr, Inorganic Chemistry, 3rd Edn. Pearson Education 2004. ❖ J. E. Huheey, E. A. Keiter, R. L. Keiter, O K Medhi, Inorganic Chemistry, Pearson 2006. ❖ J. D. Lee, Concise Inorganic Chemistry 5th edn., Wiley India Pvt. Ltd. 2008. ❖ M. Clyde Day, and J. Selbin Theoretical inorganic chemistry 2nd Edn. Reinhold Book Corp. 2008. ❖ 6. B. Douglas, D. Mc Daniel, J. Alexander, Concepts and models of Inorganic Chemistry 3rd edn., John Wiley. 2006. 		
UNIT 5 : Bioinorganic Chemistry (10 hours)	9	Essential and trace elements in biological systems, Myoglobin and Hemoglobin, role of myoglobin and hemoglobin in biological systems	Assignment No.3
	10	Mechanism of oxygen transport, cooperativity, Bohr effect, Phosphate effect	
	11	Cytochromes- Structure and function.	
	12	Metalloenzymes: Inhibition and poisoning of enzymes. A brief study of the following metalloenzymes and their functions. Carbonic anhydrase and Carbonic peptidase.	
	13	Cytochrome oxidase, cytochrome P450, Peroxidase, catalases, superoxide dismutase and Nitrogenase.	
	14	Role of alkali and alkaline earth metals in biological systems, Na/K pump.	

SECOND INTERNAL EXAMINATION			
Text Books		<ul style="list-style-type: none"> ❖ J. D. Lee, Concise Inorganic Chemistry 5th edn., Wiley India Pvt. Ltd.2008. ❖ R. Puri, L. R. Sharma, K. C. Kalia, Principles of Inorganic Chemistry,31st Milestone Publishers, New Delhi 2010 ❖ G. L. Meissler, D.A Tarr, Inorganic Chemistry,3rd Edn. Pearson Education, 2004. ❖ J. E. Huheey, E. A. Keiter, R. L. Keiter, O K Medhi, Inorganic Chemistry, Pearson 2006 ❖ F. A. Cotton and G. Wilkinson, Advanced Inorganic Chemistry 6th edn., John Wiley, New York 1991. 	
U N I T	15	Metal deficiency: Deficiency of Iron, Copper and Zinc	
	16	Metal toxicity. Toxicity of Copper, Iron, Calcium, Plutonium, Mercury and Cadmium.	
	17	Metals as carcinogens. Treatment of metal toxicity. Chelation therapy.	Demonstration
	18	Anti-cancer drugs – cisplatin and carboplatin.	PowerPoint presentation
Text Books		<ul style="list-style-type: none"> ❖ B. R. Puri, L. R. Sharma, K. C. Kalia, Principles of Inorganic Chemistry, Milestone Publishers, New Delhi, 2010. ❖ G. L. Meissler, D. A Tarr, Inorganic Chemistry, 3rd Edn. Pearson Education, 2004. ❖ J. E. Huheey, E. A. Keiter, R. L. Keiter, O K Medhi, Inorganic Chemistry,5th Ed. Pearson 2009. ❖ F.A.Cotton, G.Wilkinson, P. L. Gaus, Basic Inorganic Chemistry,3rd Edn,John –Wiley,1995 ❖ B. Douglas, D. Mc Daniel, J. Alexander, Concepts and models of Inorganic Chemistry 3rd edn., John Wiley. ❖ Ivano Bertini, Harry B Gray, Stephen J. Lippard, Joan Selverstone Valentine,Bioinorganic Chemistry.Viva Books Pvt Ltd. 2007. 	

JUC	No. of Session	Session Topic and Discussion Theme	Value additions	Remarks
UNIT 1 : Chemistry of d and f block elements (9 hours)	1	Module 1 : Chemistry of d and f Block Elements Different properties of d block elements		
	2	electronic configuration, oxidation state.		
	3	Valency, metallic character, colour.	Assignment No: 1	
	4	Magnetic properties, catalytic properties and ability to form complexes.		
	5	Comparison with second and third transition series.	Group Discussion	
	6	Chemistry of Lanthanides		
	7	Their properties	MOODLE- Assignment No:2	
		FIRST INTERNAL EXAMINATION		
Text Books	<ul style="list-style-type: none"> ❖ J. D. Lee, Concise Inorganic Chemistry 5th edn., Wiley India Pvt. Ltd.2008. ❖ R. Puri, L. R. Sharma, K. C. Kalia, Principles of Inorganic Chemistry,31st Milestone Publishers, New Delhi 2010 ❖ G. L. Meissler, D.A Tarr, Inorganic Chemistry,3rd Edn. Pearson Education, 2004. ❖ J. E. Huheey, E. A. Keiter, R. L. Keiter, O K Medhi, Inorganic Chemistry, Pearson 2006 ❖ F. A. Cotton and G. Wilkinson, Advanced Inorganic Chemistry 6th edn., John Wiley, New York 1991. 			
U NI	8	Discussion on CIA	Assignment No.3	

T 1 : Chemistry of d and f block	9	Lanthanide contraction, separation of lanthanides. Actinides, properties. Comparison of lanthanides and actinides		
	SECOND INTERNAL EXAMINATION			
Text Books	<ul style="list-style-type: none"> ❖ J. D. Lee, Concise Inorganic Chemistry 5th edn., Wiley India Pvt. Ltd.2008. ❖ R. Puri, L. R. Sharma, K. C. Kalia, Principles of Inorganic Chemistry,31st Milestone Publishers, New Delhi 2010 ❖ G. L. Meissler, D.A Tarr, Inorganic Chemistry,3rd Edn. Pearson Education, 2004. ❖ J. E. Huheey, E. A. Keiter, R. L. Keiter, O K Medhi, Inorganic Chemistry, Pearson 2006 ❖ F. A. Cotton and G. Wilkinson, Advanced Inorganic Chemistry 6th edn., John Wiley, New York 1991. 			
UNIT 3 : Organometallic Compounds (9 hours)	10	Definition, Classification of organometallic compounds,		
	11	Ylides, Classification on the basis of hapticity,		
	12	Naming of organometallic compounds.	Demonstration	
	13	catalytic properties, alkene hydrogenation, shift reaction,	PowerPoint presentation	
	14	Zeigler-Natta polymerization, 18 e rule,		
	15	Metal-alkene complexes, metal-alkyne complexes,		
	16	Metallocenes-Ferrocene. Zeise salt.		
	17	Preparation and structure.		
	18	Revision		
Text Books	<ul style="list-style-type: none"> ❖ B. R. Puri, L. R. Sharma, K. C. Kalia, Principles of Inorganic Chemistry,31st Edn. Milestone Publishers, New Delhi 2010 ❖ G. L. Meissler, D. A Tarr, Inorganic Chemistry,3rd Edn. Pearson Education, 2004. ❖ J. E. Huheey, E. A. Keiter, R. L. Keiter, O K Medhi, Inorganic Chemistry, Pearson 2006 ❖ R. C. Mehrothra and A. Singh, Organometallic chemistry, New age publishers. ❖ F. A. Cotton and G. Wilkinson, Advanced Inorganic Chemistry 3rd edn., John Wiley, New York.1995. ❖ A. G. Sharpe, Inorganic Chemistry, 3rd Edn. Pearson. 			

COURSE PLAN				
ACADEMIC YEAR 2016-17				
PROGRAMME	:	<i>B.Sc. Chemistry</i>	LECTURE HOURS	: 54
SEMESTER	:	5	CREDITS	: 3
SUBJECT TITLE	:	Organic Chemistry - III	SUBJECT CODE	: 15U5CRCHE06
COURSE TEACHERS	:	<i>Dr. V.S. Sebastian, Dr. Joseph T Moolayil, Dr. Franklin John, Dr. M. George</i>		
COURSE OUTCOMES (COs)	:	<p>CO1 - To learn the chemistry of nitro compounds, amines, dyes, organic polymers, soaps, detergents and organic reagents.</p> <p>CO2 - To understand and study mechanism of reactions of nitro compounds and amines</p> <p>CO3 - To have an elementary idea of chemotherapy, organic spectroscopy and photochemistry</p> <p>CO4 - <i>Ability to explore and reflect about the wide range of possibilities and applications of nuclear reactions and radio activity.</i></p> <p>CO5 - To identify organic compound using UV, IR and PMR spectroscopic techniques</p>		
	No. of Session	Session Topic and Discussion Theme	Value additions	COs
C Orga	1	Introduction Nitro compounds- nitromethane- tautomerism reduction products of nitrobenzene in acidic, neutral and alkaline media-		CO1
	2	reduction products of nitrobenzene in acidic, neutral		CO1

nic com pou nds cont aini ng Nitr ogen (10 hour s)	3	reduction products of nitrobenzene in alkaline media-	Assignment No: 1	CO1
	4	Electrolytic reduction and selective reduction of poly nitro compounds- formation of charge transfer complexes		CO1
	5	Amines- isomerism- stereochemistry of amines. Separation of a mixture of primary, secondary and tertiary amines -		CO1
	6	Structural features affecting basicity of aliphatic and aromatic amines. Quaternary amine salts as phase-transfer catalysts		CO1, CO2
	7	Comparative study of aliphatic and aromatic amines.	Assignment No:2	CO1, CO2
	8	Preparation of alkyl and arylamines (reduction of nitro compounds, nitriles),		CO1, CO2, CO3
	9	FIRST INTERNAL EXAMINATION		
	Text Books	<ol style="list-style-type: none"> 1. I. L. Finar, Organic Chemistry -, 6th Edition. Vol.- I, Pearson. (Chapters13, 22, 23, 24). 2. R. T. Morrison and R.N Boyd, 'Organic Chemistry', 6th Edition - Prentice Hall of India, (Chapter- 22,23). 3. M. K. Jain and S. C. Sharma 'Modern Organic Chemistry', 3rd Edition, Vishal Publishing Company Co. (Chapter-22). 4. K. S. Tewari and N. K. Vishnoi, 'Organic Chemistry', 3rd Edition, Vikas Publishing House (Chapter- 22,23,24). 5. B. S. Bahl, 'Advanced Organic Chemistry', S. Chand. 		
	10	Reductive amination of aldehydes and ketones Gabriel- Phthalimide reaction, Hoffmann bromamide reaction.	Assignment No.3	CO1
	11	Diazonium salts-preparation, synthetic transformations of aryldiazonium salts		CO1
	No. of Session	Session Topic and Discussion Theme	Value additions	COs
	12			CO1

	13	Azo Coupling - Mechanisms of Sandmeyer's and Gatterman reactions	Group Discussion	CO1, CO3
SECOND INTERNAL EXAMINATION				
<i>Text Books</i>	<ol style="list-style-type: none"> 1. I. L. Finar, Organic Chemistry -, 6th Edition. Vol.- I, Pearson. (Chapters13, 22, 23, 24). 2. R. T. Morrison and R.N Boyd, 'Organic Chemistry', 6th Edition - Prentice Hall of India, (Chapter- 22,23). 3. M. K. Jain and S. C. Sharma 'Modern Organic Chemistry', 3rd Edition, Vishal Publishing Company Co. (Chapter-22). 4. K. S. Tewari and N. K. Vishnoi, 'Organic Chemistry', 3rd Edition, Vikas Publishing House (Chapter- 22,23,24). 5. B. S. Bahl, 'Advanced Organic Chemistry', S. Chand. 			
Organic compounds containing Nitrogen (5 hours)	14	Schiemann and Gomberg reactions		CO1
	15	Preparation and uses of Phenyl hydrazine		CO1, CO5
	16	Diazomethane - preparation, structure and synthetic uses	Demonstration	CO1, CO5
	17	Diazoacetic ester - preparation, structure and synthetic uses	PowerPoint presentation	CO1, CO5
	18	Arndt- Eistert synthesis- mechanism	PowerPoint presentation	CO1, CO5
	19	Wolff rearrangement –mechanism		
	20	Curtius rearrangement and its mechanism.		
<i>Text Books</i>	<ol style="list-style-type: none"> 6. I. L. Finar, Organic Chemistry -, 6th Edition. Vol.- I, Pearson. (Chapters13, 22, 23, 24). 7. R. T. Morrison and R.N Boyd, 'Organic Chemistry', 6th Edition - Prentice Hall of India, (Chapter- 22,23). 8. M. K. Jain and S. C. Sharma 'Modern Organic Chemistry', 3rd Edition, Vishal Publishing Company Co. (Chapter-22). 9. K. S. Tewari and N. K. Vishnoi, 'Organic Chemistry', 3rd Edition, Vikas Publishing House (Chapter- 22,23,24). 10. B. S. Bahl, 'Advanced Organic Chemistry', S. Chand. 			



Unit II & III , IV & V	No. of Sessions	Session Topic and Discussion Theme	Value additions	COs
Dyes (3 hours)	1	Theory of colour and constitution. Classification - according to structure and method of application.		CO1
	2	Preparation and uses of Azo dye-methyl orange and Bismark brown,		CO1, CO4
	3	Preparation and uses of Triphenyl methane dye -Malachite green,		CO1
	4	Preparation and uses of Phthalein dye - phenolphthalein and fluroescen,		CO1
	5	Preparation and uses of Vat dye – indigo,		CO1, CO4
	7	Preparation and uses of Anthraquinone dye - alizarin		CO1, CO4
	1st Internal Examination			
Text Books	<ul style="list-style-type: none"> ○ I. L. Finar, Organic Chemistry, 6th Edition. Vol - I, Pearson. (Chapter-31) ○ M. K. Jain and S. C. Sharma ‘Modern Organic Chemistry’, 3rd Edition, Vishal Publishing Company Co. (Chapter-22) ○ K. S. Tewari and N. K. Vishnoi, ‘Organic Chemistry’, 3rd Edition, Vikas Publishing House (Chapter-38). ○ B. S. Bahl, ‘Advanced Organic Chemistry’, S. Chand 			
Unit V, VI & VII	No. of Sessions	Session Topic and Discussion Theme	Value additions	COs
Photo ch	10	Introduction- Photochemical versus Thermal reactions. Reactions		CO1, CO2, CO3

emical Reactions (3 hours), Organic Polymers (4hours), Soaps and Detergents (3) hours	11	Norrish reactions of acyclic Ketones.	Power Point Presentation	CO1, CO2, C03	
	12	Patterno-Buchi, Photo-Fries reactions.	Power Point Presentation	CO1, CO2, C03	
	13	Synthesis and applications of the following polymers- Polyesters- terephthalates,		CO1	
	14	Nylon 6 and Nylon 6,6, phenol formaldehyde resins, urea formaldehyde resins .	Group Discussion	CO1	
	2nd Internal Examination				
	15	Epoxy resins and polyurethanes, PVC and Teflon.		CO1	
	16	Synthetic rubbers –SBR and Nitrile rubber- structure and applications		CO1	
	17	Composition of soaps- detergent action of soap	Individual Assignment:	CO1	
	18	Synthetic detergents- - their functions – comparison between soaps and detergents-		CO1	
	19	Environmental aspects. LAS and ABS detergents			
Text Books	<ol style="list-style-type: none"> 1. I. L. Finar, Organic Chemistry, 6th Edition. Vol- I, Pearson. (p.-323) 2. M. K. Jain and S. C. Sharma, ‘Modern Organic Chemistry’, 3rd Edition, Vishal Publishing Company Co. (Chapter-22) 3. K. S. Tewari and N. K. Vishnoi ‘Organic Chemistry’, 3rd Edition, Vikas Publishing House (Chapter-36). 4. R. T. Morrison and R.N Boyd, ‘Organic Chemistry’, 6th Edition - Prentice Hall of India, (Chapter- 31) 5. Billmeyer F.W., Text book of polymer science, Jr.John Wiley and Sons, 1994. 6. Gowariker V.R., Viswanathan N.V. and JayaderSreedhar, Polymer Science, Wiley Eastern Ltd., New Delhi. 				
VSS					
Unit VIII & IX	No. of Sessions	Session Topic and Discussion Theme	Value additions	COs	

Aliphatic hydrocarbons (2 hours), Chemotherapy (5) hours Chemotherapy	1	Cycloalkanes- relative stabilities..		CO1
	2	Butadiene – structure and stability, 1,4 addition and its mechanism		CO1, CO4
	3	Drugs- introduction –classification –mode of action		CO1
	4	Elementary idea of the structure and mode of action of drugs Sulphanilamides,		CO1
	5	Elementary idea of the structure and mode of action of drugs Amphicillin and Chloramphenicol		CO1, CO4
	7	Elementary idea of the structure and application of Chloroquine, Paracetamol, Analgin and Aspirin.		CO1, CO4
	1st Internal Examination			
Text Books	<ul style="list-style-type: none"> ○ I. L. Finar, Organic Chemistry, 6th Edition. Vol - I, Pearson. (Chapter-31) ○ M. K. Jain and S. C. Sharma ‘Modern Organic Chemistry’, 3rd Edition, Vishal Publishing Company Co. (Chapter-22) ○ K. S. Tewari and N. K. Vishnoi, ‘Organic Chemistry’, 3rd Edition, Vikas Publishing House (Chapter-38). ○ B. S. Bahl, ‘Advanced Organic Chemistry’, S. Chand 			
Unit IV & V	No. of Sessions	Session Topic and Discussion Theme	Value additions	COs
Chemistry of Organic	10	Drugs in cancer therapy- Chlorambucil		CO1, CO2, CO3
	11	Analytical reagents – Tollens reagent, Fehling solution	Power Point Presentation	CO1, CO2, CO3
	12	Schiff’s reagents, Borsche’s reagent, Benedict solution-(Procedure not required.	Power Point Presentation	CO1, CO2, CO3

Reagents (4 hours), Structure elucidation (8)	13	Applications of Synthetic reagents –NBS, Lead tetra acetate, Periodic acid, OsO ₄		CO1
	14	Ozone, LDA, Raney Nickel, Selenium dioxide, DCC (elementary idea.	Group Discussion	CO1
	2nd Internal Examination			
	15	Introduction to UV, IR and NMR spectroscopy.		CO1
	16	UV, IR and NMR spectral characteristics of simple molecules such as ethylene, butadiene, benzene, acetaldehyde, acetone acetophenone, crotonaldehyde, ethanol		CO1
	17	Problems pertaining to the structure elucidation of simple organic compounds using IR and PMR spectroscopic techniques	Individual Assignment:	CO1
	18	Mass spectrometry- Introduction-EI ionisation- Determination Molecular mass by MS (Elementary idea- fragmentation study not required)		CO1
Text Books	I. L. Finar, Organic Chemistry, 6 th Edition. Vol- I, Pearson. (p.-323) M. K. Jain and S. C. Sharma, 'Modern Organic Chemistry', 3 rd Edition, Vishal Publishing Company Co. (Chapter-22) K. S. Tewari and N. K. Vishnoi 'Organic Chemistry', 3 rd Edition, Vikas Publishing House (Chapter-36). R. T. Morrison and R.N Boyd, 'Organic Chemistry', 6 th Edition - Prentice Hall of India, (Chapter- 31)			

DEPARTMENT OF CHEMISTRY, SACRED HEART COLLEGE (AUTONOMOUS), THEVARA	
COURSE PLAN : ACADEMIC YEAR 2016 - 2017	
PROGRAMME : <i>B.Sc. Chemistry</i>	SEMESTER : 5
LECTURE HOURS : 36	CREDITS : 2
SUBJECT TITLE : <i>States of Matter</i>	SUBJECT CODE : <i>CH5B03</i>
COURSE TEACHERS	: Dr. K. B. Jose (KBJ), Dr. Thommachan Xavier (TX) & Dr. Ignatious Abraham (IGA)
COURSE OUTCOMES (COs)	: CO1 - <i>Understand the intermolecular forces and dynamics of molecules in gases and liquids</i> CO2 - <i>Describe the structure of solids and defects in crystals.</i> CO3 - <i>Understand and explain surface phenomena.</i>
Instructional Hours	: <i>Tuesday : Period 1 (9:30 to 10:30 am) - KBJ</i> <i>Wednesday : Period 2 (10:30 to 11:30 am) - IGA</i> <i>Thursday : Period 2 (10:30 to 11:30 pm) - TX</i>

IGNATIUS ABRAHAM			
Unit I : GASES			
Session s	Session Topic and Discussion Theme	Value additions	COs
1	Kinetic molecular model of gases		CO1
2	Pressure of an ideal gas, derivation of gas laws		CO1
3	Maxwell's distribution of velocities – molecular velocities (average, root mean square and most probable velocities)		CO1
4	Collision diameter, mean free path		CO1
5	Viscosity of gases – temperature and pressure dependence. Relation between mean free path and coefficient of viscosity.	Power Point Presentation	CO1
6	Barometric distribution law		CO1
7	Law of equipartition of energy		CO1
8	Degrees of freedom and molecular basis of heat capacities.		CO1
9	Real gases: compressibility factor z		CO1
1 st Internal Examination			
10	van der Waals equation of state – derivation and application in explaining real gas behaviour.		CO1
11	Virial equation of state	<i>Assignment :</i> Synthetic Applications of active methylene compounds	CO1
12	Van der Waals equation expressed in virial form – calculation of Boyle temperature		CO1
13	Isotherms of real gases		CO1
14	Continuity of states. Critical phenomena.		CO1
15	Liquefaction of gases		CO1

2nd Internal Examination**Unit III : SYMMETRY**

16	Symmetry of molecules-symmetry elements and symmetry operations – centre of symmetry, plane of symmetry, proper and improper axes of symmetry,	Power Point Presentation	CO2
17	Combination of symmetry elements, molecular point groups, Schoenflies symbol,	<i>Assignment</i>	CO2
18	Crystallographic point groups		CO2

References:

1. B. R. Puri, L. R. Sharma, M. S. Pathania, Elements of Physical chemistry, Vishal Pub. Co. Jalandhar, Chapters 1,2
2. K. L. Kapoor, A Textbook of Physical chemistry, Volumes 1, Macmillan India Ltd Chapter 1
3. P. Atkins and J Paula, The elements of Physical chemistry, 7th edn., Oxford University Press, Chapter 1
4. F. A. Alberty and R J Silby, Physical Chemistry, 3rd Edn, John Wiley, Chapter 17

THOMMACHAN XAVIER**Unit III : LIQUIDS**

Session s	Session Topic and Discussion Theme	Value additions	COs
1	Liquid State - introduction		CO1
2	Intermolecular forces in liquids	Power point presentation	CO1
3	Viscosity – Factors affecting viscosity		CO1
4	the viscometer method	Power Point Presentation:	CO1
5	surface tension		CO1
6	Determination of surface tension		CO1
7	Structure of liquids.		CO1
8	Unusual behaviour of water		CO1

Unit III : LIQUID CRYSTALS AND ADSORPTION

9	Classification of liquid crystals	<i>Assignment</i>	CO2
1st Internal Examination			
10	Structure of nematic phases		CO2
11	Structure of cholestric phases		CO2
12	Adsorption – types		CO3
13	Adsorption of gases by solids		CO3
14	Factors influencing adsorption		CO3
15	Freundlich adsorption isotherm	Power Point Presentation	CO3
2nd Internal Examination			
16	Langmuir adsorption isotherm		CO3
17	The BET theory		CO3
18	Use of BET equation for the determination of surface area.		CO3

References:

1. K. L. Kapoor, A Textbook of Physical chemistry, Volume 1, Macmillan India Ltd Chapters 2,3
2. P. Atkins and J. Paula, The elements of Physical chemistry, 7th edn., Oxford University Press, Chapter 23
3. A. McQuarrie, J. D. Simon, Physical Chemistry – A molecular Approach, Viva Books Pvt. Ltd, Chapter 29
4. B. R. Puri, L. R. Sharma, M. S. Pathania, Elements of Physical Chemistry, Vishal Publishing Co, Chapter

K B JOSE

Unit III : SOLID STATE

Session s	Session Topic and Discussion Theme	Value additions	COs
1	The nature of the solid state	Power Point Presentation	CO2
2	Anisotropy- the law of constancy of interfacial angles		CO2
3	Law of rational indices - Miller indices.		CO3
4	Seven crystal systems and fourteen Bravais lattices.	Assignment	CO2
5	X-ray diffraction, Bragg's law		CO2
6	Detailed study of simple, face centred and body centred cubic systems	Models	CO2
7	Bragg's x-ray diffractometer method	Power Point	CO2
8	Powder pattern method.		CO2
9	Analysis of powder diffraction patterns of NaCl and KCl	Assignment	CO2
1st Internal Examination			
10	Density of cubic crystals, identification of cubic crystal from crystallographic data.	Assignment	CO2
11	Close packing of spheres, ccp and hcp arrangements.		CO2
12	Structure of ionic compounds of the type AX - NaCl	Power Point	CO2
13	Structure of ionic compounds of the type AX - CsCl, ZnS	Power Point	CO2
14	Structure of ionic compounds of the type AX ₂ - (CaF ₂ , Na ₂ O)	Power Point	CO2
15	Defects in crystals – stoichiometric and non-stoichiometric defects	Assignment	CO2
2nd Internal Examination			
16	Extrinsic and intrinsic defects.		CO2
17	Electrical conductivity, semiconductors, n-type, p-type		CO2
18	Superconductivity – an introduction	Power Point	CO2

References:

1. Peter Sykes, A Guide book to Mechanism in Organic Chemistry: 6th Edition, Pearson Education.
2. P. S. Kalsi 'Organic Reactions and their Mechanisms' New Age International Publishers.
3. K.S. Tewari and N.K. Vishnoi 'Organic Chemistry', 3rd Edition, Vikas Publishing House.
4. M. K. Jain and S.C. Sharma 'Modern Organic Chemistry', 3rd Edition, Vishal Publishing Company Co.
5. R. T. Morrison and R. N. Boyd, 'Organic Chemistry', 6th Edition - Prentice Hall of India,
6. I. L. Finar, Organic Chemistry, 6th Edition. Vol.- I, Pearson

COURSE PLAN

ACADEMIC YEAR 2016-17

PROGRAMME	:	<i>B.Sc. Chemistry</i>	LECTURE HOURS	:	36
SEMESTER	:	5	CREDITS	:	2
SUBJECT TITLE	:	Physical Chemistry - II	SUBJECT CODE	:	-
COURSE TEACHERS	:	<i>Dr.Jinu George, Senju Devassykutty</i>			
COURSE OBJECTIVES	:	<ul style="list-style-type: none"> • Understand the basics of spectroscopy. • Explain the fundamental principles of rotational, vibrational, Raman, electronic, NMR and mass spectroscopic techniques. • Describe the fundamentals of photochemistry and optical properties. 			
Instructional Hours	:	<i>2 hours per week</i>			

	No. of Sessions	Session Topic and Discussion Theme	Value additions	
4. Raman spectroscopy	1	➤ Electronic Spectroscopy: Introduction		
	2	Electronic spectrum: concept of potential energy curves for bonding and anti-bonding molecular orbitals		https://youtu.be/MW4PwJxxyt0
	3	➤ electronic transition, the Frank-Condon principle,	ICT	
	4	dissociation energy. Polyatomic molecules – qualitative description of σ , π and n- molecular orbitals		
	5	➤ Polyatomic molecules- energy levels and the respective transitions		

5. Electronic spectroscopy	7	NMR Spectroscopy : NMR spectroscopy: basic principles of NMR spectroscopy	ICT	
	1st Internal Examination			
	8	Nuclear spin, Larmor precession. Proton magnetic resonance (¹ H NMR or PMR)		https://youtu.be/t5stvnKNXbg
	9	Nuclear shielding and deshielding, chemical shift and molecular structure. Spin-spin splitting and coupling constant.		
Text Books	References 1. R. Puri, L. R. Sharma, M. S. Pathania, ' <i>Elements of Physical Chemistry</i> ', Vishal Pub. Co., 2. K. J. Laidler, John H. Meiser, ' <i>Physical Chemistry</i> ', 2 nd edn..			
	No. of Sessions	Session Topic and Discussion Theme	Value additions	
6. NMR Spectroscopy 7. Mass Spectroscopy	10	➤ First order spectra – interpretation of PMR spectra of simple organic molecules		
	11	➤ First order spectra examples- ethyl bromide, ethanol, acetaldehyde, ethyl acetate, toluene, acetophenone.	Power Point Presentation	
	12	➤ Mass Spectroscopy: Mass spectrometry: Basic principle-ionization	Power Point Presentation	
	13	➤ Fragmentation, separation of ions and representation of the spectrum		
	14	Application in molecular mass determination.	Group Discussion	
	2nd Internal Examination			
	15	➤ Photochemistry: Interaction of radiation with matter: Laws of photochemistry – Grothus-Draper law, Stark-Einstein law, examples of photochemical reactions.		

	16	Beer law and Beer-Lambert's law. Jablonsky diagram, qualitative description of fluorescence		
	17	Quantum yield, primary and secondary processes. Basic concepts of photosensitized reactions – photosynthesis, dissociation of hydrogen molecule, isomerization of 2-butene, and chemiluminescence.		
	18	Optical properties - optical activity, molar refraction.		
<i>Text Books</i>	References <ul style="list-style-type: none"> • K. K. Sharma, L R Sharma, 'A Text Book of Physical Chemistry', Vikas Publishing house. • S. Negi, S. C. Anand, 'A Textbook of Physical Chemistry', Second Edition, New Age International (P) limited, publishers. 			

	No. of Sessions	Session Topic and Discussion Theme	Value additions	Courses/ Text book/Web URL
Unit .1. Introduction to spectroscopy and	1	1. Introduction to spectroscopy Introduction: electromagnetic radiation, regions of the spectrum,		Presentation on: Molecules and Molecular Spectroscopy - UAF 12/17/2017
	2	interaction of electromagnetic radiation with molecules, various types of molecular spectroscopic techniques,		
	3	Born-Oppenheimer approximation.		

3. Vibrational spectroscopy	4	2. Rotational spectroscopy Introduction to Rotational spectrum:	ICT		
	5	diatomic molecules, energy levels of a rigid rotator, selection rules, determination of bond length.			
	7	3. Vibrational spectroscopy			
	1st Internal Examination				
	8	Vibrational spectrum: the simple harmonic oscillator			
	9	– energy levels, force constant, selection rules.			
<i>Text Books</i>	References 1. R. Puri, L. R. Sharma, M. S. Pathania, ' <i>Elements of Physical Chemistry</i> ', Vishal Pub. Co., 2. K. J. Laidler, John H. Meiser, ' <i>Physical Chemistry</i> ', 2 nd edn..				
2. Rotational spectroscopy	No. of Sessions	Session Topic and Discussion Theme	Value additions		
2. Rotational spectroscopy 4. Raman	10	Anharmonic oscillator			
	11	– pure vibrational spectra of diatomic molecules,	Power Point Presentation		
	12	selection rules, fundamental frequencies, overtones,	Power Point Presentation		
	13	hot bands. Degrees of freedom for polyatomic molecules,			
	14	revision	Group Discussion		

spectroscopy	2nd Internal Examination			
	15	concept of group frequencies –		Classical Mechanics with a Bang! - URL is "https://modphys.hosted.uark.edu/markup/CMwBang_Units_Detail_2017.html"
	16	– pure vibrational spectra of diatomic molecules,		
	17	frequencies of common functional groups in organic compounds.		
18	Raman spectrum: quantum theory of Raman Effect (elementary idea), concept of polarizability, qualitative treatment of pure rotational and pure vibrational Raman spectra of diatomic molecules, selection rules, rule of mutual exclusion.			
Text Books	References <ul style="list-style-type: none"> • K. K. Sharma, L R Sharma, ‘<i>A Text Book of Physical Chemistry</i>’, Vikas Publishing house. • S. Negi, S. C. Anand, ‘<i>A Textbook of Physical Chemistry</i>’, Second Edition, New Age International (P) limited, publishers. 			

SACRED HEART COLLEGE(AUTONOMOUS), THEVARA

DEPARTMENT OF CHEMISTRY

COURSE PLAN

ACADEMIC YEAR 2016-2017

PROGRAMME	Open course	SEMESTER	5
COURSE CODE AND TITLE	15U5OCHE1 - Chemistry in Everyday life	CREDIT	3
HOURS/SEM	72		
FACULTY NAME	Dr. Joseph T Moolayil (JTM), Dr. Grace Thomas (GT), Dr. Ramakrishnan S (RKS), Dr. Abi T G (ATG)		

COURSE OBJECTIVES	
1	Understand chemistry of Food additives and Flavours
2	Understand chemistry of Soaps
3	Understand chemistry of synthetic detergent
4	Understand chemistry of Cosmetics
5	Understand chemistry of Plastics, Paper Dyes
6	Understand chemistry of Drugs
7	Understand chemistry of Chemistry and Agriculture

Dr. Abi T. G.				
Plastics, Paper Dyes				(14 Hrs)
Sl.No	Session	Topic	Method of Teaching	Remarks
1	1	Plastics in everyday life	Chalk & Board	
2	2	Brief idea of polymerization-	Chalk & Board	
3	3	Thermoplastic and thermosetting polymers.	Chalk & Board	
4	4	Use of PET, HDPE, PVC, LDPE, PP, ABS.	Chalk & Board	
5	5	Use of PET, HDPE, PVC, LDPE, PP, ABS.	Chalk & Board	
First Internal Examination				
6	6	Biodegradable plastics	Chalk & Board	
7	7	Environmental hazards of plastics	Chalk & Board	
8	8	News print paper, writing paper, paper	Chalk & Board	

		boards, cardboards.		
9	9	Organic materials, wood, cotton, jute and coir.	Chalk & Board	
10	10	International recycling codes, and symbols for identification.	Chalk & Board	
11	11	Natural and synthetic dyes (basic idea only).	Chalk & Board	
12	12	Recycling of plastics.	Chalk & Board	
<i>Second internal Examination</i>				
13	13	Revision	Chalk & Board	
14	14	Revision	Chalk & Board	
Synthetic Detergents				(3 Hrs)
15	15	Enzymes used in commercial detergents	Chalk & Board	
16	16	Environmental hazards.	Chalk & Board	
17	17	Revision	Chalk & Board	

Dr. Joseph T Moolayil				
Cosmetics				(12 Hrs)
1	1	Cosmetics- Introduction,	Chalk & Board	
2	2	classification	Chalk & Board	
3	3	bathing oils	Chalk & Board	
4	4	toilet powder,	Chalk & Board	
5	5	dental cosmetics	Chalk & Board	
6	6	shaving cream	Chalk & Board	
<i>First Internal Examination</i>				
7	7	shampoo, hair dyes	Chalk & Board	
8	8	face creams	Chalk & Board	
9	9	skin products	Chalk & Board	
10	10	General formulation of each type.	Chalk & Board	
11	11	Toxicology of cosmetics	Chalk & Board	
12	12	revision	Chalk & Board	
Soaps				(7 Hrs)

13	13	Soaps – Introduction	Chalk & Board	
<i>Second Internal Examination</i>				
14	14	Detergent action of soap.	Chalk & Board	
15	15	Toilet soap, bathing bars	Chalk & Board	
16	16	Washing soaps, liquid soap manufacture-. Significance of acidity and alkalinity.	Chalk & Board	
17	17	Additives, fillers and flavours	Chalk & Board	
18	18	Significance of acidity and alkalinity	Chalk & Board	
19	19	Revision	Chalk & Board	

Dr. Ramakrishnan S

Food additives and Flavours

(12 Hrs)

1	1	Functional food additives	Chalk & Board	
2	2	adulteration	Chalk & Board	
3	3	food laws	Chalk & Board	
4	4	food laws	Chalk & Board	
5	5	Food colours - permitted and non – permitted-	Chalk & Board	
6	6	Food colours: Toxicology.	Chalk & Board	

First Internal Examination

7	7	Flavours – natural and synthetic-	Chalk & Board	
8	8	Flavours – Toxicology	Chalk & Board	
9	9	Other functional additives	Chalk & Board	
10	10	Soft drinks- formulation	Chalk & Board	
11	11	Health drinks	Chalk & Board	
12	12	Revision	Chalk & Board	

Synthetic Detergents

(6 Hrs)

13	13	Detergents- Introduction,	Chalk & Board	
<i>Second Internal Examination</i>				
14	14	detergent action	Chalk & Board	
15	15	types of detergents-cationic, anionic, amphiphilic detergents.	Chalk & Board	
16	16	Common detergent chemicals.	Chalk & Board	

17	17	Additives, excipients colours and flavours.	Chalk & Board	
18	18	Revision	Chalk & Board	

Dr. Grace Thomas				
Drugs			(6 Hrs)	
1	1	Chemotherapy	Chalk & Board	
2	2	- types of drugs- analgesics,	Chalk & Board	
3	3	- types of drugs- antipyretics, antihistamines	Chalk & Board	
4	4	- types of drugs- antacids tranquilizers, sedatives	Chalk & Board	
5	5	- types of drugs: antibiotics	Chalk & Board	
6	6	- types of drugs- antifertility drugs.	Chalk & Board	
<i>First Internal Examination</i>				
Chemistry and Agriculture			(12 Hrs)	
7	7	Fertilizers- natural, synthetic, mixed	Chalk & Board	
8	8	NPK fertilizers.	Chalk & Board	
9	9	Excessive use of fertilizers and its impact on the environment.	Chalk & Board	
10	10	Bio fertilizers. Plant growth hormones.	Chalk & Board	
11	11	Pesticides- Classification-insecticides, herbicides, fungicides.	Chalk & Board	
12	12	Excessive use of pesticides – environmental hazards.	Chalk & Board	
13	13	Excessive use of pesticides – environmental hazards.	Chalk & Board	
<i>Second Internal Examination</i>				
14	14	Bio pesticides.	Chalk & Board	
15	15	Antiseptics and Disinfectants	Chalk & Board	
16	16	Disinfectants-Oils - vegetable oils, mineral oil	Chalk & Board	
17	17	essential oil-Sugars, artificial sugars	Chalk & Board	
18	18	Revision	Chalk & Board	

ASSIGNMENTS AND SEMINARS

Sl No	Module	Topic	Nature of Assignment	Remarks
1	1	Excessive use of pesticides – environmental hazards.	Case studies in short	
2	2	Ingredients of any 2 cosmetics	written	

Reference books

1	P. Coultate, Food- The Chemistry of its components. Royal Society of Chemistry, London(Paper back)
2	Shashi Chowls, Engineering Chemistry, Danpat Rai Publication.
3	B.K. Sharma. Industrial Chemistry
4	CNR Rao- Understanding chemistry, Universities Press.
5	Puri and Sharma. Advanced Organic Chemistry.
6	Brown, Insect control by chemicals
7	A. K. De, Environmental Chemistry, New age International Ltd.
8	S. S. Dara, A Textbook of Environmental chemistry and pollution control, S.Chand & Company Ltd
9	Tisdale, S.L., Nelson, W.L. and Beaton, J. D. Soil Fertility and Fertilizers, Macmillian Publishing Company, New York, 1990.
10	Buchel, K.H. Chemistry of Pesticides, John Wiley & Sons, New York, 1983
11	P.C Pall, K. Goel, R.K Gupta, Insecticides, pesticides and agrobased industries.
12	Gowariker V.R., Viswanathan N.V. and Jayader Sreedhar, Polymer Science, Wiley Eastern Ltd., New Delhi.
13	I.I Singh, V.K Kapoor, Organic Pharmaceutical Chemistry

COURSE PLAN

ACADEMIC YEAR 2016-17

PROGRAMME	:	<i>B.Sc. Chemistry</i>	LECTURE HOURS	:	<i>54</i>
SEMESTER	:	<i>6</i>	CREDITS	:	<i>3</i>
SUBJECT TITLE	:	<i>Inorganic Chemistry II</i>	SUBJECT CODE	:	<i>15U6CRCHE09</i>
COURSE TEACHERS	:	<i>Dr. Joseph John (JJ), Mr. Midhun Dominic C D (MD), Ms. June Cyriac (JUC)</i>			
Objectives	:	<p>To understand the principle of metallurgical processes, the preparation and uses of inorganic polymers, importance of non-aqueous chemistry, metal carbonyls, the structure of solids and the general characteristics of p-block elements.</p> <p>To understand the importance of our environment and its protection.</p>			
Instructional Hours	:	<i>3 hours per week</i>			

JJ	No. of Session	Session Topic and Discussion Theme	Value additions	Remarks
UNIT 1 : Metallurgy	1	Methods of concentration of ores- Gravity, magnetic and electrostatic separations, Froth flotation and leaching		
	2	Calcination and Roasting. Reduction to free metal- smelting and electrometallurgy,		
	3	Hydrometallurgy. Goldschmidt Thermite Process.	Assignment No: 1	
	4	Refining of metals- electrolytic, ion exchange, zone refining, vapour		

(9 hours)		phase refining and oxidative refining.		
	5	Thermodynamics of the oxidation of metals to metal oxides - Ellingham diagrams.	Group Discussion	
	6	Extractive metallurgy of U, Ti		
	7	Extractive metallurgy of Th and Ni.	MOODLE- Assignment No:2	
	8	Revision		
	9	FIRST INTERNAL EXAMINATION		
Text Books	<ul style="list-style-type: none"> ❖ B. R. Puri, L. R. Sharma, K. C. Kalia, Principles of Inorganic Chemistry, 31st Ed. Milestone Publishers, New Delhi 2010. ❖ S. Prakash, G. D. Tuli, S. K. Basu and R. D. Madan, Advanced Inorganic Chemistry, 5th edn, 2012, Volume I, S Chand. ❖ A. Cottrell, An introduction to metallurgy, 2nd edn., University press. 1990. 			
UNIT 6: Structure of Inorganic Solids (9 hours)	10	Close packing of spheres, ccp and hcp arrangements.	Assignment No.3	
	11	Interstitial sites in close packing, Tetrahedral, Octahedral sites. Radius ratio, Limiting radius ratio for trigonal, tetrahedral and octahedral sites. (only values).		
	12	Use of limiting radius ratio in the structural determination of ionic crystals. Structure of ionic crystals of NaCl		
	13	Structure of ionic crystals of CsCl, ZnS		
	14	Defects in crystals – stoichiometric and non-stoichiometric defects		
SECOND INTERNAL EXAMINATION				

Text Books	<ul style="list-style-type: none"> ❖ B. R. Puri, L. R. Sharma, K C Kalia, Principles of Inorganic Chemistry, 31st Edn. Milestone Publishers, New Delhi, 2010. ❖ J. D. Lee, Concise Inorganic Chemistry 5th edn., Blackwell Science, London, 2008. ❖ J. E. Huheey, E. A. Keiter, R. L. Keiter, O K Medhi, Inorganic Chemistry, 4th edn., Pearson 2006 ❖ D. F. Shriver and P.W. Atkins, Inorganic Chemistry, , 3rd edn., Oxford University Press. ❖ G.L. Meissler, D.A Tarr, Inorganic Chemistry, Pearson Education ❖ B. Douglas, D. Mc Daniel, J. Alexander, Concepts and models of Inorganic Chemistry 3rd edn., John Wiley. ❖ A. R. West, Solid State Chemistry and its applications, John Wiley. 		
UNIT 6: Structure of Inorganic Solids (9 hours)	15	Consequences of defects. extrinsic and intrinsic defects	
	16	Impurity defects.	Demonstration
	17	Semiconductors, n-type, p-type, Superconductivity – an introduction.	PowerPoint presentation
	18	Revision	
Text Books	<ul style="list-style-type: none"> ❖ B. R. Puri, L. R. Sharma, K C Kalia, Principles of Inorganic Chemistry, 31st Edn. Milestone Publishers, New Delhi, 2010. ❖ J. D. Lee, Concise Inorganic Chemistry 5th edn., Blackwell Science, London, 2008. ❖ J. E. Huheey, E. A. Keiter, R. L. Keiter, O K Medhi, Inorganic Chemistry, 4th edn., Pearson 2006 ❖ D. F. Shriver and P.W. Atkins, Inorganic Chemistry, , 3rd edn., Oxford University Press. ❖ G.L. Meissler, D.A Tarr, Inorganic Chemistry, Pearson Education ❖ B. Douglas, D. Mc Daniel, J. Alexander, Concepts and models of Inorganic Chemistry 3rd edn., John Wiley. ❖ A. R. West, Solid State Chemistry and its applications, John Wiley. 		

MD	No. of Session	Session Topic and Discussion Theme	Value additions	Remarks
UNIT 2 : Metal Carbonyls and Metal Clusters (9 hours)	1	Introduction to metal carbonyls and metal clusters		
	2	Preparation and properties of mononuclear carbonyls.		
	3	Structures of Mo(CO) ₆ , Fe(CO) ₅ and Ni(CO) ₄ .	Assignment No: 1	
	4	Polynuclear carbonyls, bridged carbonyls and bonding in carbonyls.		
	5	Metal clusters - carbonyl and halide clusters	Group Discussion	
	6	Low nuclearity carbonyl clusters and high nuclearity carbonyl clusters,		
	7	Electron counting schemes for Rh ₆ (CO) ₁₆ and [Os ₆ (CO) ₁₈] ²⁻	MOODLE- Assignment No:2	
	8	Metal only clusters (Zintl ions). Quadruple bond – structure of Re ₂ Cl ₈ ²⁻ .		
	9	Revision		
FIRST INTERNAL EXAMINATION				
Text Books	<ul style="list-style-type: none"> ❖ B. R. Puri, L. R. Sharma, K. C. Kalia, Principles of Inorganic Chemistry, 31st Edn. Milestone Publishers, New Delhi 2010. ❖ G. L. Meissler, D. A Tarr, Inorganic Chemistry, 3rd Edn. Pearson Education 2004. ❖ J. E. Huheey, E. A. Keiter, R. L. Keiter, O K Medhi, Inorganic Chemistry, Pearson 2006. ❖ J. D. Lee, Concise Inorganic Chemistry 5th edn., Wiley India Pvt. Ltd. 2008. ❖ M. Clyde Day, and J. Selbin Theoretical inorganic chemistry 2nd Edn. Reinhold Book Corp. 2008. ❖ 6. B. Douglas, D. Mc Daniel, J. Alexander, Concepts and models of Inorganic Chemistry 3rd edn., John Wiley. 2006. 			

UNIT 5 Compounds of s and p block elements	10	Introduction to s and p block elements		
	11	Macrocyclic ligands:- crown ethers and cryptands, Macrocyclic effect	Assignment No.3	
	12	Alkali metal complexes with crown ethers and cryptands, their applications.		
	13	Boron hydrides – diborane (preparation, properties and bonding)		
	14	B ₅ H ₉ , B ₄ H ₁₀ (structure only). Closo carboranes, boron nitride		
	15	Borazine, boric acid and Peroxy acids of sulphur.		
	16	Oxides and oxy acids of halogens (structure only), superacids,		
	17	Interhalogen compounds, pseudohalogens, electropositive iodine, (structure only).		
	18	Fluorocarbons. Fluorides, oxides and oxy fluorides of xenon (structure only).		
		SECOND INTERNAL EXAMINATION		
Text Books		<ul style="list-style-type: none"> ❖ J. D. Lee, Concise Inorganic Chemistry 5th edn., Blackwell Science, London,2008. ❖ B. R. Puri, L. R. Sharma, K C Kalia, Principles of Inorganic Chemistry, 31st Edn.Milestone Publishers, New Delhi,2010. ❖ J. E. Huheey, E. A. Keiter, R. L. Keiter, O K Medhi, Inorganic Chemistry, 4th edn., Pearson 2006. ❖ D. F. Shriver and P.W. Atkins, Inorganic Chemistry, 3rd edn., Oxford University Press, 2006. ❖ M. N. Greenwood and A. Earnshaw, Chemistry of the elements 2nd edn, Butterworth, 1997. 		

JUC	No. of Session	Session Topic and Discussion Theme	Value additions	Remarks
<i>UNIT 3 : Inorganic Polymers (6 hours)</i>	1	Inorganic polymers – general properties, comparison with organic polymers		
	2	Glass transition temperature. Sulphur based polymers – polymeric sulphur nitride and chalcogenic glasses (preparation)		
	3	Sulphur based polymers – polymeric sulphur nitride and chalcogenic glasses (properties and uses).	Assignment No: 1	
	4	Phosphorus based polymers – polyphosphazenes and polyphosphates.		
	5	Silicon based polymers – silicones and silicone rubber (preparation)	Group Discussion	
	6	Silicon based polymers – silicones and silicone rubber (properties and uses).		
	FIRST INTERNAL EXAMINATION			
<i>Text Books</i>	<ul style="list-style-type: none"> ❖ B. R. Puri, L. R. Sharma, K. C. Kalia, Principles of Inorganic Chemistry, 31st Edn. Milestone Publishers, New Delhi, 2010. ❖ S. Prakash, G. D. Tuli, S. K. Basu and R. D. Madan, Advanced Inorganic Chemistry, Volume I, S Chand. ❖ J. E. Huheey, E. A. Keiter, R. L. Keiter, O K Medhi, Inorganic Chemistry, 4th edn., Pearson 2006 			
<i>UNIT 4 : Non-Aqueous Solvents (4</i>	7	Classification of solvents, characteristics of solvents	Assignment No.3	
	8	Reactions in liquid ammonia, Alkali metal solution in liquid ammonia, their important properties.		
	9	Liquid sulphur dioxide (acid base, amphoteric, solvation, oxidation – reduction, complex formation)		

<i>hours</i>	10	Liquid HF (acid base, amphoteric, solvation, oxidation – reduction, complex formation)		
	SECOND INTERNAL EXAMINATION			
<i>Text Books</i>	<ul style="list-style-type: none"> ❖ B. R. Puri, L. R. Sharma, K. C. Kalia, Principles of Inorganic Chemistry, 31st Milestone Publishers, New Delhi, 2010 ❖ J. E. Huheey, E. A. Keiter, R. L. Keiter, O K Medhi, Inorganic Chemistry, 4th edn., Pearson 2006. ❖ M. Clyde Day, and J. Selbin, Theoretical Inorganic Chemistry, 2nd Edn. Reinhold Book Corp. 2008. 			
<i>UNIT 7 : Water Quality Parameters (7 hours)</i>	11	Introduction to water quality parameters		
	12	Standards for drinking water-Determination of turbidity (nephelometric method)		
	13	Determination of pH-determination of total dissolved solids	Demonstration	
	14	Total hardness-total alkalinity-acidity - chloride	PowerPoint presentation	
	15	Determination of dissolved oxygen (DO), BOD		
	16	Determination of COD. Estimation of coliform count.		
	17	Revision		
<i>Text Books</i>	<ul style="list-style-type: none"> ❖ B.K Sharma Environmental Chemistry, 12th Edn., Goel Publishing House, 2011. ❖ B. R. Puri, L. R. Sharma, K C Kalia, Principles of Inorganic Chemistry, 31st Edn. Milestone Publishers, New Delhi, 2010. ❖ A.K Dee, Environmental Chemistry, 3rd Edn., New Age International Pvt. Ltd., 1996. ❖ Sodhi. G.S., Concepts of Environmental Chemistry, Narsa Publication House, 2009. ❖ Sindhu. P. S., Environmental Chemistry, New Age International Pvt. Ltd., 2011. ❖ Balaram Pani, Environment Chemistry, I. K. International Publishing House Ltd., 2007. ❖ Thomas G Spiro, Chemistry of Environment, Prentice Hall of India., 2006. ❖ Raghavan Nambiar., Environmental Studies, Scitech Publications (India) Pvt. Ltd., 2008. 			

COURSE PLAN

ACADEMIC YEAR 2016-17

PROGRAMME	:	<i>B.Sc. Chemistry</i>	LECTURE HOURS	:	54
SEMESTER	:	6	CREDITS	:	3
SUBJECT TITLE	:	<i>Organic Chemistry-IV</i>	SUBJECT CODE	:	19U6CRCHE10
COURSE TEACHERS	:	<i>V.S. Sebastian (VSS), Franklin J (FJ), Joseph T Moolayil (JTM), M. George (MG)</i>			

FJ	No. of Session	Session Topic and Discussion Theme	Value additions	
UNIT 1 : Natural Products (12 hours)	1	Natural Products - Terpenoids		
	2	Isoprene rule. Structure elucidation of citral and geraniol		
	3	Structure elucidation of geraniol	Assignment No: 1	
	4	Alkaloids - general methods of isolation		
	5	Alkaloids-classification – structure elucidation		
	6	Synthesis of coniine		
	7	Synthesis of piperine	Assignment No:2	
	8	Synthesis of nicotine.		
	9	Vitamins – classification- structure (elementary idea) of vitamin A, C and B1, B2, B6		

		FIRST INTERNAL EXAMINATION			
	Text Books	<ul style="list-style-type: none"> ❖ I. L. Finar, <i>Organic Chemistry</i> - Volume I & II - Pearson Education. ❖ M. K. Jain and S. C. Sharma '<i>Modern Organic Chemistry</i>', 3rd Edition, Vishal Publishing Company Co . ❖ K.S. Tewari and N.K. Vishnoi, '<i>Organic Chemistry</i>', 3rd Edition, Vikas Publishing House. 			
	10	Lipids – biological functions – oils and fats – common fatty acids			
	11	Extraction and refining- hydrogenation –			
	12	Rancidity- identification of oils and fats			
	13	Revision-Natural products, alkaloids			
		SECOND INTERNAL EXAMINATION			
	Text Books	<ul style="list-style-type: none"> ❖ L. Finar, <i>Organic Chemistry</i> - Volume I & II - Pearson Education. ❖ M. K. Jain and S. C. Sharma '<i>Modern Organic Chemistry</i>', 3rd Edition, Vishal Publishing Company Co. ❖ K.S. Tewari and N.K. Vishnoi, '<i>Organic Chemistry</i>', 3rd Edition, Vikas Publishing House. 			
JTM					
	Unit II	No. of Sessions	Session Topic and Discussion Theme	Value additions	
Carbohydrates (14 Hours)		1	Classification - constitution of glucose and fructose.		
		2	Reactions of glucose-osazone formation		
		3	Reactions of fructose - osazone formation.		
		4	Reactions of glucose and fructose - Mutarotation and its mechanism.		
		5	Epimerisation		
		6	Configuration of monosaccharides		
			Ist Internal Examination		
		7	Cyclic structure. Pyranose and furanose forms		
		8	Determination of ring size.	Power Point Presentation	
	9	Determination of ring size. Haworth projection formula.			

	10	Chain lengthening and chain shortening of aldoses.		
	11	Inter conversion of aldoses and ketoses.		
	12	Disaccharides - reactions and structure of sucrose and maltose. Ring structure		
	2nd Internal Examination			
	13	Structure and properties of starch and cellulose (elementary idea).		
	14	Industrial applications of cellulose.		
	Text Books	<ul style="list-style-type: none"> ❖ I. L. Finar, <i>Organic Chemistry</i> - Volume I & II - Pearson Education. ❖ M. K. Jain and S. C. Sharma '<i>Modern Organic Chemistry</i>', 3rd Edition, Vishal Publishing Company Co. ❖ K.S. Tewari and N.K. Vishnoi, '<i>Organic Chemistry</i>', 3rd Edition, Vikas Publishing House. ❖ R. T. Morrison and R.N. Boyd, '<i>Organic Chemistry</i>', 6th Edition - Prentice Hall of India 		
MG				
Unit III	No. of Sessions	Session Topic and Discussion Theme	Value additions	
Heterocyclic Compounds (10 Hours)	1	Aromaticity of heterocyclic compounds.		
	2	Preparation, properties and uses of furan	Power Point Presentation	
	3	Preparation, properties and uses of pyrrole	Power Point Presentation	
	4	Preparation, properties and uses of thiophene.		
	5	Synthesis and reactions of pyridine	Group Discussion	
	6	Synthesis and reactions of piperidine -		

	7	comparative study of basicity of pyrrole, pyridine and piperidine with amines.		
	8	Synthesis and reactions of quinoline, isoquinoline and indole with special reference to Skraup synthesis		
	9	Bischler, Napieralskii and Fisher indole synthesis		
	1st Internal Examination			
	10			
Unit VI	No. of Sessions	Session Topic and Discussion Theme	Value additions	
Steroids (3 Hours)	1	Introduction – Diels hydrocarbon-	Individual Assignment:	
	2	Structure and functions of cholesterol.		
	2nd Internal Examination			
	3	Elementary idea of HDL, LDL, Vitamin D		
Text Books	<ul style="list-style-type: none"> ❖ I. L. Finar, Organic Chemistry - Volume I & II - Pearson Education. ❖ M. K. Jain and S. C. Sharma ‘Modern Organic Chemistry’, 3rd Edition, Vishal Publishing Company Co. ❖ K.S. Tewari and N.K. Vishnoi, ‘Organic Chemistry’, 3rd Edition, Vikas Publishing House. ❖ R. T. Morrison and R.N. Boyd, ‘Organic Chemistry’, 6th Edition - Prentice Hall of India 			

VSS				
Unit IV	No. of Sessions	Session Topic and Discussion Theme	Value additions	
Amino acids and Proteins	1	Amino acids- classification,		
	2	Zwitter ion. Peptide-		
	3	Solution phase peptide synthesis.	Power Point Presentation	
	4	Classification of proteins based on physical and chemical		

Compounds (10 Hours)		properties and on physiological functions.			
	5	Primary secondary tertiary and quaternary structure of proteins	Group Discussion		
	6	Helical and sheet structures (<i>elementary treatment only</i>). Nucleic acids. Types of nucleic acids			
	7	Nucleic acids. Types of nucleic acids,			
	8	RNA and DNA,			
	9	polynucleotide chain components			
	Ist Internal Examination				
	10	Green Fluorescent Proteins (<i>elementary idea</i>)			
	Unit V	No. of Sessions	Session Topic and Discussion Theme	Value additions	
Enzymes (3 Hours)	1	Nomenclature and classification of enzymes (<i>based on substrate</i>).	Individual Assignment:		
	2	Chemical nature of enzymes. Mechanism of enzyme action.			
	2nd Internal Examination				
	3	Substrate specificity of enzymes. Enzyme inhibition.			

<i>Text Books</i>	<ul style="list-style-type: none">❖ I. L. Finar, Organic Chemistry - Volume I & II - Pearson Education.❖ M. K. Jain and S. C. Sharma 'Modern Organic Chemistry', 3rd Edition, Vishal Publishing Company Co.❖ K.S. Tewari and N.K. Vishnoi, 'Organic Chemistry', 3rd Edition, Vikas Publishing House.❖ R. T. Morrison and R.N. Boyd, 'Organic Chemistry', 6th Edition - Prentice Hall of India <ul style="list-style-type: none">❖ en.wikipedia.org/wiki/Green_fluorescent_protein❖ www.scholarpedia.org/article/fluorescent_protein❖ www.conncoll.edu/ccacad/zimmer/GFP-ww/timeline.html❖ www.gonda.ucla.edu/bri_core/gfp.htm
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DEPARTMENT OF CHEMISTRY, SACRED HEART COLLEGE (AUTONOMOUS), THEVARA			
COURSE PLAN : ACADEMIC YEAR 2016 - 2017			
PROGRAMME	: <i>B.Sc. Chemistry</i>	SEMESTER	: 6
LECTURE HOURS	: 54	CREDITS	: 3
SUBJECT TITLE	: <i>Equilibrium and Kinetics</i>	SUBJECT CODE	: <i>CHE6B03</i>
COURSE TEACHERS	:	Dr. Ignatious Abraham (IGA), Dr. K. B. Jose (KBJ) & Senju Devassykutty (SD)	
COURSE OUTCOMES (COs)	:	CO1 - <i>Understand the basics of thermodynamics.</i> CO2 - <i>Understand the applicability of the laws of thermodynamics to various physical and chemical processes.</i> CO3 - <i>Describe the phase diagrams of one- and two-component systems</i> CO4 - <i>Understand the basic principles of chemical kinetics.</i>	
Instructional Hours	:	<i>Monday : Period 3 (9:30 to 10:30 am) - SD</i> <i>Tuesday : Period 1 (9:30 to 10:30 am) - KBJ</i> <i>Friday : Period 2 (10:30 to 11:30 am) - IGA</i>	

IGNATIOUS ABRAHAM			
Unit I : CLASSICAL THERMODYNAMICS			
Session s	Session Topic and Discussion Theme	Value additions	COs
1	Introduction to Thermodynamics: Definition of thermodynamic terms, intensive and extensive properties		CO1
2	Path and state functions, exact and inexact differentials		CO1
3	Reversible and irreversible processes,		CO1
4	Spontaneous and non-spontaneous processes, internal energy, work and heat		CO1
5	Zeroth law of thermodynamics	Power Point Presentation	CO1
6	First law of thermodynamics: Statement and mathematical expression		CO1
7	Enthalpy, heat capacity, Cp and Cv relation in ideal gas systems,		CO1
8	Change in thermodynamic properties of an ideal gas during isothermal reversible / irreversible processes.		CO2
9	Change in thermodynamic properties of an ideal gas during adiabatic, reversible / irreversible processes.		CO2
1 st Internal Examination			
10	Joule-Thomson experiment,		CO2
11	Joule-Thomson coefficient μ_{JT} , inversion temperature	<i>Assignment :</i> Synthetic Applications of active methylene compounds	CO2
12	Second law of Thermodynamics: Limitations of first law – statements of second law,		CO1

13	Carnot's cycle – efficiency of heat engines, Carnot theorem.		CO1
14	Entropy – entropy change for various reversible/irreversible processes,		CO1
15	Change in entropy of an ideal gas with pressure, volume and temperature.		CO1

2nd Internal Examination

Unit III : SYMMETRY

16	Third law of thermodynamics-statement and significance.	Power Point Presentation	CO2
17	Helmholtz energy and Gibbs energy	<i>Assignment</i>	CO1
18	Variation of Gibbs energy with T and P		CO1

References:

1. R. P. Rastogi, R. R. Misra, An Introduction to Chemical Thermodynamics, 6th edn., Vikas Pub. Pvt. Ltd.
2. K. L. Kapoor, A Textbook of Physical chemistry, Volumes 3, Macmillan India Ltd. Chapters 3, 5, 6.
3. P. Atkins and J Paula, The elements of Physical chemistry, 7th edn., Oxford University Press, Chapter 8.
4. B. R. Puri, L. R. Sharma, M. S. Pathania, Elements of Physical chemistry, Vishal Pub. Co. Jalandher.
5. J. Rajaram and J. C. Kuriakose, Thermodynamics, ShobanLal Nagin Chand & Co (1986).
6. H. Kuhn and H. D. Fosterling, Principles of Physical chemistry, John Wiley.
7. W. J. Moore, Basic Physical Chemistry, Orient Longman.

SENJU DEVASSYKUTTY

Unit II : PHASE EQUILIBRIA

Session s	Session Topic and Discussion Theme	Value additions	COs
1	The phase rule, equilibrium between phases – conditions.		CO3
2	One component system – water system	Power point presentation	CO3
3	One component system - sulphur system		CO3
4	Two component systems – solid-liquid equilibrium – simple eutectic,	Power Point Presentation:	CO3
5	Lead- silver system		CO3
6	Formation of compounds with congruent melting point ferric chloride- water system,		CO3
7	Formation of compounds with incongruent melting point sodium sulphate- water system.		CO3

Unit I : THERMOCHEMISTRY

8	Enthalpies of formation and combustion		CO1
9	Enthalpies of neutralization, solution and hydration	<i>Assignment</i>	CO1

1st Internal Examination

10	Relation between heats of reactions at constant volume and constant pressure.		CO2
11	Variation of heats of reaction with temperature – Kirchoff's equation		CO2
12	Hess's law and its application.		CO1
13	Criteria for reversible and irreversible processes.		CO1
14	Gibbs-Helmholtz equation.		CO1
15	Clausius - Clapeyron equation, applications.	Power Point Presentation	CO1

2 nd Internal Examination			
16	Partial molar properties – chemical potential,		CO2
17	Gibbs-Duhem equation		CO2
18	Chemical potential in a system of ideal gases, concept of activity.		CO2

References:

1. R. P. Rastogi, R. R. Misra, An Introduction to Chemical Thermodynamics, 6th edn., Vikas Pub. Pvt. Ltd.
2. K. L. Kapoor, A Textbook of Physical chemistry, Volumes 3, Macmillan India Ltd. Chapters 3, 5, 6.
3. P. Atkins and J Paula, The elements of Physical chemistry, 7th edn., Oxford University Press, Chapter 8.
4. B. R. Puri, L. R. Sharma, M. S. Pathania, Elements of Physical chemistry, Vishal Pub. Co. Jalandher.
5. J. Rajaram and J. C. Kuriakose, Thermodynamics, ShobanLal Nagin Chand & Co (1986).
6. H. Kuhn and H. D. Fosterling, Principles of Physical chemistry, John Wiley.
7. W. J. Moore, Basic Physical Chemistry, Orient Longman.

K B JOSE

Unit III : SOLID STATE

Session s	Session Topic and Discussion Theme	Value additions	COs
1	Rate of reaction, rate equation, order and molecularity of reactions	Power Point Presentation	CO4
2	Integrated rate expressions for first and second order reactions.		CO4
3	Zero order reactions, pseudo-order reactions, half-life.		CO4
4	Theories of chemical kinetics: effect of temperature on the rate of reaction	<i>Assignment</i>	CO4
5	Arrhenius equation, concept of activation energy		CO4
6	Collision theory, transition state theory.	Models	CO4
7	Thermodynamic parameters for activation – Eyring equation (no derivation needed),	Power Point	CO4
8	Enthalpy and entropy of activation.		CO4
9	Theory of unimolecular reactions – Lindemann theory.	Assignment	CO4

1st Internal Examination

10	Kinetics of complex (composite) reactions: Opposing reactions, consecutive reactions, and parallel (simultaneous) reactions.	<i>Assignment</i>	CO4
11	Chain reactions – steady state treatment, hydrogen bromine reaction.		CO4
12	Catalysis: Homogeneous catalysis,	Power Point	CO4
13	Enzyme catalysis – Michaelis-Menten equation (no derivation needed).	Power Point	CO4
14	Heterogeneous catalysis – surface catalysis, uni and bi molecular reactions on surface.	Power Point	CO4
15	Elementary idea about autocatalysis.	<i>Assignment</i>	CO4

2nd Internal Examination

Unit I : Chemical Equilibrium

16	Chemical equilibrium: conditions for chemical equilibrium.		CO2
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17	van't Hoff reaction isotherm, relation between K_c and $K_x - K_p$		CO2
18	Temperature dependence of K_p – van't Hoff equation	Power Point	CO2

References:

1. J. Rajaram and J. C. Kuriakose, Thermodynamics, ShobanLal Nagin Chand & Co (1986).
2. H. Kuhn and H. D. Fosterling, Principles of Physical chemistry, John Wiley.
3. W. J. Moore, Basic Physical Chemistry, Orient Longman.
4. B. R. Puri, L. R. Sharma, M. S. Pathania, Elements of Physical Chemistry, Vishal Pub. Co. Jalandhar.
5. D. A. McQuarrie, J. D. Simon, Physical Chemistry – A molecular Approach Viva Books Pvt. Ltd.
6. K. L. Kapoor, A Textbook of Physical Chemistry, Volumes 4, Macmillan India Ltd.
7. K. K. Sharma, L. K. Sharma, A Textbook of Physical Chemistry, 4th edn, Vikas publishing House.

	No. of Session	Session Topic and Discussion Theme	Value additions	WEB url/ADDITIONAL RESOURCES
UNIT 1 : Ionic equilibrium	1	Introduction-concepts of acids and bases	Power point	
	2	relative strength of acid-base pairs, influence of solvents	Chalk & Board	
	3	Classification of acids and bases as hard and soft acids and bases. Pearson's HSAB concept, applications,.	Chalk & Board	
	4	Dissociation constants – acids, bases, and polyprotic acids.	Chalk & Board	
	5	Ostwald's dilution law. Ionic product of water – pH.	Chalk & Board	
	6	Buffer solutions – mechanism of buffer action,	Assignment No:1	
	7	Henderson equation. Hydrolysis of salts – hydrolysis constant, degree of hydrolysis, pH of salt solutions.(contd derivation)	Chalk & Board	
	8	Acid-base indicators, theories, determination of pH by indicators, solubility product principle – applications.	Power point	
	9	FIRST INTERNAL EXAMINATION		
Text Books	<ul style="list-style-type: none"> ❖ K. L. Kapoor, '<i>A Textbook of Physical Chemistry</i>', Volumes 1, Macmillan India Ltd. ❖ B. R. Puri, L. R. Sharma, M. S. Pathania, '<i>Elements of Physical Chemistry</i>', Vishal Pub. Co. Jalandhar. ❖ I. N. Levine, '<i>Physical Chemistry</i>', Tata Mc Graw Hill. 			
UNI T 4:	10	Introduction, optical activity	Power point	

Electrical and Magnetic properties	11	molar refraction, dielectric constant,	Chalk & Board	
	(10 hours) No. of Session	Session Topic and Discussion Theme		Value additions
	12	dipole moment, magnetic properties of molecules,	Power point	
	13	nuclear paramagnetism,	Group Discussion	
SECOND INTERNAL EXAMINATION				
Text Books	<ul style="list-style-type: none"> ❖ K. J. Laidler and J. M. Meiser, <i>Physical Chemistry</i>, 3rd Edition, Houghton Mifflin Comp., New York, International Edition (1999). ❖ Barrow, G.M. <i>Physical Chemistry</i>, Tata McGraw-Hill (2007). ❖ Castellan, G.W. <i>Physical Chemistry</i>, 4th Ed. Narosa (2004). 			
UNIT 4: Electrical and Magnetic properties	14	NMR spectrometer	Assignment No:2	
	15	the chemical shift	Power point	
	16	electron spin resonance (ESR).	Demonstration	
	17	numericals	PowerPoint presentation	
	18	revision	PowerPoint presentation	
Text Books	<ul style="list-style-type: none"> ❖ F A Alberty and R J Silby, <i>Physical Chemistry</i>, John Wiley. ❖ P. W. Atkins, <i>The elements of Physical chemistry</i>, 8th edn, Oxford University Press. ❖ S. H. Marron and J. B. Lando, <i>Fundamentals of Physical Chemistry</i>, Macmillan Ltd. 			
COURSE PLAN				

ACADEMIC YEAR 2016-17				
PROGRAMME	:	<i>B.Sc. Chemistry</i>	LECTURE HOURS	: 54
SEMESTER	:	6	CREDITS	: 3
SUBJECT TITLE	:	<i>Physical Chemistry-IV</i>	SUBJECT CODE	: 15U6CRCHE12
COURSE TEACHERS	:	<i>Dr Jinu George (JG), Dr.Thommachan Xavier, Dr. K B Jose</i>		
COURSE OBJECTIVES	:	<ul style="list-style-type: none"> - Ability to understand the concept of acids, bases and pH of solutions. - Ability to find out pH of various solutions using appropriate equations. - Ability to understand the electrical and magnetic properties of substances. - Ability to understand the conductance properties of electrolytes and their applications. - Ability to understand electrochemical equivalent and types of cells and its working so that student can work in an electrochemical laboratory 		
Instructional Hours	:	<i>3 hours per week</i>		
	No. of Session	Session Topic and Discussion Theme	Value additions	
<i>U NI</i>	1	Introduction - Faraday's laws of electrolysis	Demonstration	

T 2 and UNIT 2 : Electrical Conductance	2	electrochemical equivalent, and chemical equivalent	Demonstration	
	3	Electrolytic conductivity, molar conductivity - Variation of molar conductivity with concentration.	Assignment No: 3	
	4	Kohlrausch's law – applications.	Group discussion	
	5	Ionic mobility – relation with ion conductivity, influence of temperature on ion conductivity,	Chalk & board	
	6	ion conductivity and viscosity – Walden's rule	Chalk & board	
	7	Influence of dielectric constant of solvent on ion conductivity. Abnormal ion conductivity of hydrogen and hydroxyl ions.	Chalk & board	
	8	Discharge of ions during electrolysis – Hittorf's theoretical device.	Discussion	
		FIRST INTERNAL EXAMINATION		
Text Books	<ul style="list-style-type: none"> ❖ Mahan, B.H. University Chemistry, 3rd Ed. Narosa (1998). ❖ Glasstone S, An Introduction to Electrochemistry, East-West Press (Pvt.) Ltd. (2006). ❖ Gurdeep Raj, Advanced Physical Chemistry, Goel publishing house. 			
	9	Transport Numbers – determination by Hittorf's method and moving boundary method.	Assignment No.3	
	10	Debye-Hückel theory of strong electrolytes	Chalk & board	
	11	The concept of ionic atmosphere, Asymmetry and electrophoretic effect.	Chalk & board	
	12	Debye- Hückel-Onsager equation (<i>no derivation</i>)	Group Discussion	
	SECOND INTERNAL EXAMINATION			

Text Books	<ul style="list-style-type: none"> ❖ Mahan, B.H. University Chemistry, 3rd Ed. Narosa (1998). ❖ Glasstone S, An Introduction to Electrochemistry, East-West Press (Pvt.) Ltd. (2006). ❖ Gurdeep Raj, Advanced Physical Chemistry, Goel publishing house. 			
Electrical Conductance	13	Activity, mean ionic activity and mean ionic activity coefficients of electrolytes.	PowerPoint presentation	
	14	Ionic strength of a solution, Debye-Hückel limiting law (<i>no derivation</i>)	PowerPoint presentation	
	15	Applications of conductance measurements	Demonstration	
	16	Determinations of degree of dissociation of weak electrolytes, ionic product of water	PowerPoint presentation	
	17	Solubility of sparingly soluble salts	PowerPoint presentation	
	18	conductometric titrations.	PowerPoint presentation	
Text Books	<ul style="list-style-type: none"> ❖ Mahan, B.H. University Chemistry, 3rd Ed. Narosa (1998). ❖ Glasstone S, An Introduction to Electrochemistry, East-West Press (Pvt.) Ltd. (2006). ❖ Gurdeep Raj, Advanced Physical Chemistry, Goel publishing house. 			
Unit IV	No. of Sessions	Session Topic and Discussion Theme	Value additions	
UNIT 3: Elect	1	Introduction – Electrochemical Cells and Electrolytic cells, Galvanic cells	PowerPoint presentation	
	2	Characteristics of reversible cells. Reversible electrodes – different	PowerPoint presentation	

<i>romotive Force</i>		types			
	3	Reference electrodes – Standard Hydrogen Electrode, Calomel electrode, electrode potential – electrochemical series.	Group discussions		
	4	Representation of cells – e.m.f of cell, electrode reactions and cell reactions.	Group discussions		
	5	Thermodynamics of reversible cells and reversible electrodes – Determination of ΔG , ΔH and ΔS of cell reaction.	PowerPoint presentation		
	7	E.M.F and equilibrium constant of cell reaction	PowerPoint presentation		
	1st Internal Examination				
	8	Effect of electrolyte concentration on electrode potential and e.m.f - Derivation of Nernst equation.	Chalk & board		
9	Concentration cells – electrode concentration cell and electrolyte concentration cells	Power Point Presentation			
<i>Text Books</i>	<ul style="list-style-type: none"> ❖ Glasstone S, An Introduction to Electrochemistry, East-West Press (Pvt.) Ltd. (2006). ❖ Gurdeep Raj, Advanced Physical Chemistry, Goel publishing house. ❖ F A Alberty and R J Silby, Physical Chemistry, John Wiley. ❖ P. W. Atkins, The elements of Physical chemistry, 8th edn, Oxford University Press. 				
<i>Unit IV</i>	No. of Sessions	Session Topic and Discussion Theme	Value additions		
	10	Types of electrolyte concentration cells – with transference and without transference	Power Point Presentation		
	11	Liquid junction potential. Fuel cells – the hydrogen-oxygen fuel cell.	Power Point Presentation		
	12	Applications of e.m.f measurements – determination of solubility product	Power Point Presentation		
	13	determination of pH using hydrogen electrode	Chalk & board		

	14	quinhydrone electrode and glass electrode	Group Discussion	
	2nd Internal Examination			
	15	Potentiometric titrations - Redox indicators.	Power Point Presentation	
	16	Irreversible electrode processes - overvoltage.	Power Point Presentation	
	17	Corrosion of metals - forms of corrosion	Individual Assignment	
	18	Corrosion monitoring and prevention methods.	Group discussions	
<i>Text Books</i>	<ul style="list-style-type: none"> ❖ Glasstone S, An Introduction to Electrochemistry, East-West Press (Pvt.) Ltd. (2006). ❖ Gurdeep Raj, Advanced Physical Chemistry, Goel publishing house. ❖ F A Alberty and R J Silby, Physical Chemistry, John Wiley. ❖ P. W. Atkins, The elements of Physical chemistry, 8thedn, Oxford University Press. 			

COURSE PLAN**ACADEMIC YEAR 2016-17**

PROGRAMME	:	<i>B.Sc. Chemistry</i>	LECTURE HOURS	:	72
SEMESTER	:	6	CREDITS	:	4
SUBJECT TITLE	:	Advances in Chemistry	SUBJECT CODE	:	15U6ELCHE1
COURSE TEACHERS	:	Dr. Abi T G (ATG), Dr. Grace Thomas (GT), Dr Ramakrishnan S (RKS), Dr Joseph John (JJ)			
Course Objectives	:	<ol style="list-style-type: none">1. <i>Knowing the advance topics in organic, physical, inorganic and theoretical chemistry</i>2. <i>Understanding the chemistry of nanomaterials, Industrially important materials and learn about its applications in the current scenario</i>3. <i>Understand, the basic principles of Modern Analytical Techniques in Chemistry</i>4. <i>Understanding the basics principles of Green Chemistry, Supramolecular Chemistry and explore the different schemes in retrosynthetic analysis</i>5. <i>Understanding the basics of biophysical chemistry and knowing the applications of biopolymers</i>6. <i>Understanding the different computational chemistry methods and its scope in modern chemistry.</i>			

Teacher I – ATG Module Taken : Unit 1.1: Advanced Topics in Inorganic Chemistry – Nanomaterials (Total – 9 Hrs) Instructional Hours : <i>One Hour Per Week</i>			
No. of Session	Session Topic and Discussion Theme	Method of Teaching	Remarks/Student Assignmets
1	Nanomaterials	Conventional Lecture	<p style="text-align: center;">Seminar Assignment to Students on the Topic Applications of Carbon nanotubes</p>
2	Synthesis of nanomaterials – chemical precipitation,	Lecture with ICT- PPTs	
3	Mechano-chemical method and micro emulsion method,		
4	Reduction technique, chemical vapour deposition and sol-gel method (brief study).		
5	Properties and applications of fullerenes,		
6	Quantum dots.		
7	Carbon nanotubes.		
8	Applications of nano materials - nano composites –		
9	Nano medicines		
References 1. V. S. Muraleedharan and A. Subramanian, Nano science and nanotechnology, Ane Books Pvt. Ltd. New Delhi, 2009 2. T. Pradeep, Nano; The Essentials, Mc Graw-Hill education, New Delhi, 2006.			
First Internal Test			
Teacher I – ATG Module Taken : Unit 3.2 Advanced Topics in Physical Chemistry-Introduction to Computational Chemistry Instructional Hours : <i>One Hour Per Week</i>			
10	Scope of Computational chemistry. Building of 3D molecular structures using computer	Lecture with ICT- PPTs	

	softwares. Coordinate formats		
11	Z-matrix, Cartesian coordinate and PDB format. Z-matrix of simple molecules H ₂ O, CO ₂ & NH ₃ .		
12	Introduction to Common computational and visualization softwares.		
13	Brief introduction to Hartree Fock, ab initio, semi empirical methods.		
14	DFT and molecular mechanics methods.		
15	. Basis sets, STO & GTO basis sets.		
Second Internal Test			
16	Potential energy surface. Local and Global minima. Single point energy calculations and Geometry optimizations	Lecture with ICT- PPTs Problem Solving ICT- Computational Chemistry Softwares	Give Assignments on Computational Chemistry Calculations Using Gaussian Software
17	Format of input and output files in Computational Chemistry Calculations.		
18	(Single point and Optimization Calculations in simple molecules such as molecules H ₂ O, CO ₂ & NH ₃ using suitable software package.		
Reference Text Books			
<ol style="list-style-type: none"> 1. T Clark , Hand book of Computational Chemistry, Wiley, New York. 2. W J Hehre, A J Shusterman, W W Huang, A laboratory book of computational organic chemistry, Wave function Inc. 3. A. R. Leach, Molecular Modeling, Longman. 4. D. C. Young, Computational Chemistry, A practical guide to applying techniques to real world problems, John Wiley 5. F. Jensen, Introduction to Computational Chemistry, John Wiley. 			

6. Guy H. Grant and W. Graham Richards, "Computational Chemistry", OCP(29)
 7. Christopher J. Cramer, John Wiley, "Essentials of Computational Chemistry".

Teacher 2 – JJ Module Taken : Unit 1.2: Advanced Topics in Inorganic Chemistry – Industrially important materials
(Total – 8 Hrs) Instructional Hours : *One Hour Per Week*

No. of Session	Session Topic and Discussion Theme	Method of Teaching	Remarks/Student Assignmets
1	Refractory materials.	1. Conventional Lecture -Chalk & Board, 2. Seminar Presentation from Students	Assignment for Class Note Preparation
2	Carbides, nitrides, borides.		
3	Graphite and graphite oxide,		
4	Intercalation compounds of alkali metals,		
5	Carbon monofluoride,		
6	Intercalation compounds of graphite with metal Halides, glass,		
7	Silicates, zeolites,		
8	Ultramarines and ceramics.		
First Internal Test			

Teacher 2 – JJ Module Taken : Unit 1.3 Advanced Topics in Inorganic Chemistry – Modern Analytical Techniques
Instructional Hours : *One Hour Per Week*

9	Thermo analytical methods:	1. Conventional Lecture -Chalk	
10	Principle of Thermo gravimetry,		
11	TGA of calcium oxalate monohydrate and Differential thermal analysis		

12	Differential scanning calorimetry. Applications.	& Board, 2. Seminar Presentation from Students	
13	Colorimetry: Principle, Beer's law, Lambert's law		
14	Absorption coefficient, transmittance, opacity.		
15	Absorbance, optical density, molar absorption coefficient.		
16	Principle of estimation of iron, chromium and ammonia.		
Second Internal Test			
Reference Text Books			
<ol style="list-style-type: none"> 1. H.H Willard, L.L. Merritt, J.A. Dean, F.A Settle, Instrumental methods of Analysis, CBS Publishers And Distributors, Delhi, 1996. 2. G.H Jeffery, J. Basset, J. Mendham, R.C Denny, Vogel's Textbook of Quantitative Chemical Analysis, ELBS 5th edn. 1996. 			

Teacher III – GT Module Taken : Unit 2.1 Advanced topics in Organic Chemistry- Supramolecular Chemistry (Total – 6 Hrs) Instructional Hours : <i>One Hour Per Week</i>			
No. of Session	Session Topic and Discussion Theme	Method of Teaching	Remarks/Student Assignmets
1	Introduction to Supramolecular Chemistry	Conventional Lecture	
2	Molecular Recognition	Lecture with ICT- PPTs	
3	Host-guest interactions.		
4	Types of non-covalent interactions.		
5	Importance of molecular recognition in DNA		
6	Importance of molecular recognition in protein structure		
First Internal Test			
References			
1. Helena Dodzuik, Introduction to Supramolecular Chemistry Springer (chapter 1,2)			
2. J. M. Lehn, Supramolecular Chemistry, VCH			
3. H. Vogtle, Supramolecular Chemistry, Wiley			
4. En.wikipedia.org/wiki/Supramolecular Chemistry			
First Internal Test			
Teacher III – GT Module Taken : Unit 2.2 Advanced topics in Organic Chemistry- Retrosynthetic analysis (Total – 6 Hrs) Instructional Hours : <i>One Hour Per Week</i>			
7	Retrosynthetic analysis and disconnection approach.	Conventional Lecture and Power Point Presentation	Seminar Presentation by
8	Basic principles and terminology.		

9	Retrosynthetic analysis of simple cyclic and acyclic alkenes,	Lecture with ICT- PPTs	Students
10	Retrosynthetic analysis of alcohols		
11	Retrosynthetic analysis of carbonyl compounds.		
12	Simple problems of retro synthesis of the above compounds		
References			
<ol style="list-style-type: none"> 1. J. Clayden, N. Greeves, S. Warren and P. Wothers, Organic Chemistry, Oxford University Press, 2001.(Chapter 30) 2. Paula Yurkanis Bruice, Organic Chemistry, 2002, (3rd Edition) (chapter 29) 3. S. Warren, Organic Synthesis, The disconnection Approach, John Wiley & Sons, 2004. 4. E. J. Corey, X-M. Cheng (1995). <i>The Logic of Chemical Synthesis</i>. New York: Wiley. ISBN 0-471-11594-0. 5. W. Carruthers, 'Modern Methods in Organic Synthesis', Cambridge University Press., 1986. 			
Second Internal Test			
13	Introduction to Green Chemistry	Lecture with ICT- PPTs	Seminar Presentation by Students
14	Need for green chemistry		
15	Twelve principles of green chemistry		
16	Examples of Green Chemistry Processes		
17	Green polymer.		
18	Polylactic acid (PLA)		
Reference Text Books			
<ol style="list-style-type: none"> 1. V. K. Ahluwalia, Green Chemistry, <i>Ane Books</i> India. 2. Anastas, P. T.; Warner, J. C. Green Chemistry: Theory and Practice Oxford University Press: New York, 1998, p.30. By permission 			

of Oxford University Press.

3. Arends I., R. Sheldon, U. Hanefeld, Green Chemistry and Catalysis, 2007 WILEY-VCH, Verlag GmbH & Co. KGaA, Weinheim.

Teacher IV – RKS Module Taken : Unit 2.4: Advanced Topics in Organic Chemistry – Advances in Polymers (Total – 9 Hrs)
Instructional Hours : One Hour Per Week

No. of Session	Session Topic and Discussion Theme	Method of Teaching	Remarks/Student Assignmets
1	Biopolymers -	Conventional Lecture	
2	biomaterials	1. Conventional Lecture -Chalk & Board,	
3	Polymers in medical field.		
4	High temperature ploymers		
5	Fire-resistant polymers.		
6	Silicones,		
7	Conducting polymers-		
8	Carbon fibers.		
9	General discussion about the biopolymers		
First Internal Test			

Teacher IV – RKS Module Taken : Unit 3.1 Advanced Topics in Physical Chemistry-Biophysical Chemistry
(Total – 10 Hrs) Instructional Hours : One Hour Per Week

10	Protein structure; Amino acids.		
11	Primary, secondary and tertiary structure;		

12	Protein folding. Significance of Van der Waals force, hydrogen bond and hydrophobic interactions.	Lecture with ICT- PPTs	Assignment as Problems Related to Henderson and Hasselbalch equation
13	Acid-Base equilibrium: .Protonation and deprotonation reactions.		
14	Biological significance of pH; Properties of proteins with emphasis on isoelectric pH..		
15	Henderson and Hasselbalch equation. Titration curves of amino acids & pK values,		
Second Internal Test			
16	Buffers & Stability of their pH	Lecture with ICT- PPTs	
17	Thermodynamics and Kinetics. Standard free energy change in biochemical reactions, exergonic,.		
18	Hydrolysis of ATP. Chemical potential. Oxidation/reduction reactions and bioenergetics.		
19	Enzyme catalysis. Michael Menton kinetics		
Reference Text Books			
<ol style="list-style-type: none"> 1. Principles of Biochemistry- Albert L. Lehninger CBS Publishers & Distributors. 2. Biochemistry – Lubert Stryer Freeman International Edition. 3. Biochemistry - Voet and Voet, John Wiley and Sons. 4. Physical Chemistry for Life Sciences, Peter Atkins and Julio de Paula, 2006, Oxford Press 5. Narayanan, P (2000) Essentials of Biophysics, New Age Int. Pub. New Delhi. 6. Roy R.N. (1999) A Text Book of Biophysics New Central Book Agency. 7. Introduction to Biophysical chemistry, R. Bruce Martin, McGraw-Hill, NY, 1964. 8. Physical Chemistry with applications to Biological systems, Ramond Chnag, Mc Millan publishing Co.inc, New York 1977. 9. Principles of Physical Biochemistry 2nd Edition, K.E. van Holde, W.C. Johnson, P.S. Ho, 			

Pearson Prentice Hall, ISBN 0-13-046427-9.

SACRED HEART COLLEGE(AUTONOMOUS), THEVARA

DEPARTMENT OF CHEMISTRY

COURSE PLAN

ACADEMIC YEAR 2016-2017

PROGRAM ME	BSc PHYSICS	SEMESTER	1
COURSE CODE AND TITLE	15U1PCHE1, GENERAL CHEMISTRY	CREDIT	2
HOURS/ SEM	36		
FACULTY NAME	Dr. Ramakrishnan S, Mr.Senju Devassykutty		

COURSE OUTCOMES (COs)	
1	Describe different models of atomic structure.
2	Define acids and bases and explain the concept of equilibrium.
3	Explain the fundamentals of nuclear chemistry and analytical chemistry.
4	Understand basics of thermodynamics

Dr. Ramakrishnan S				
MODULE I : Atomic Structure (9hrs)				
Sl.No	Session	Topic	Method of Teaching	COs
1	1	Introduction: Atoms, Planck's quantum Theory, Photoelectric effect	Chalk & Board	1
2	2	Postulates of bohr's theory, Energy levels in atom	Chalk & Board	1
3	3	origin of hydrogen spectrum	Chalk & Board	1
4	4	Sommerfeld's extension of Bohr's Theory	Chalk & Board	1
5	5	Shortcomings of Bohr Theory	Chalk & Board	1
First Internal Examination				
6	6	Dual nature of matter and radiation. Derivation of de Broglie equation	Chalk & Board	1
7	7	Wave nature of electron and quantisation of angular momentum, Heisenberg's uncertainty principle	Chalk & Board	1
8	8	Concept of orbital, Quantum numbers,	Chalk & Board	1

		shapes of orbitals		
9	9	Electronic configuration of atoms - Aufbau principle, Hund's rule of maximum multiplicity, Pauli's exclusion principle	Chalk & Board	1
MODULE II: Concept of Equilibrium: (12 hrs)				
10	10	Acids and bases	Chalk & Board	2
11	11	theories of acids and bases	Chalk & Board	2
12	12	ionic product of water, introductory idea of pH, pOH.	Chalk & Board	2
13	13	Strengths of acids and bases, K_a and K_b , pK_a and pK_b	Chalk & Board	2
14	14	buffer solution, Henderson equation	Chalk & Board	2
15	15	hydrolysis of salt, solubility	Chalk & Board	2
Second Internal Examination				
16	16	solubility product, application	Chalk & Board	2
17	17	Common ion effect, application.	Chalk & Board	2
Mr.Senju Devassykutty				
MODULE III: Nuclear Chemistry : (6 hrs)				
1	1	Stability of Nucleus	Chalk & Board	3
2	2	natural radioactivity, induced radioactivity	Chalk & Board	3
3	3	fertile and fissile isotopes, units of radioactivity.	Chalk & Board	3
4	4	Nuclear Reactions: fission and fusion, chain reactions	Chalk & Board	3
5	5	disposal of nuclear wastes	Chalk & Board	3
6	6	applications of radioactivity	Chalk & Board	3
MODULE IV: Analytical Chemistry- Basic principles: (5hrs)				
7	7	Concentration terms- molality, molarity, normality, weight percentage, ppm, and millimoles.	Chalk & Board	3
First Internal Examination				
8	8	Titrimetric method of analysis: General principle, types of titrations, requirements for titrimetric analysis.	Chalk & Board	3
9	9	Primary and secondary standards, criteria	Chalk & Board	3

		for primary standards, preparation of standard solutions, standardization of solutions		
10	10	Errors	Chalk & Board	3
11	11	Minimization of errors	Chalk & Board	3
MODULE V: Laws of Thermodynamics : (8 hrs)				
12	12	System and Surrounding	Chalk & Board	4
13	13	First law of Thermodynamics: Internal energy, Significance of internal energy change, enthalpy	Chalk & Board	4
14	14	Second law of Thermodynamics: free energy, Entropy and Spontaneity,	Chalk & Board	4
Second internal Examination				
15	15	Statement of second law based on entropy	Chalk & Board	4
16	16	Entropy change in Phase transitions	Chalk & Board	4
17	17	entropy of fusion, entropy of vaporization, entropy of sublimation	Chalk & Board	4
18	18	The concept of Gibbs's free energy- Physical significance of free energy, conditions for equilibrium & spontaneity based on ΔG values.	Chalk & Board	4
19	19	Effect of temperature on spontaneity of Reaction. Third law of thermodynamics	Chalk & Board	4

TEXTBOOKS AND REFERENCES	
1	P. L. Soni, Inorganic Chemistry.
2	C. N. R. Rao, University General Chemistry, Macmillan.
3	R. A. Day Junior, A.L. Underwood, Quantitative Analysis, 5th edn. Prentice Hall of India Pvt. Ltd. New Delhi, 1988.
4	R. Gopalan, Analytical Chemistry, S. Chand and Co., New Delhi.
5	B. R. Puri, L. R. Sharma, M.S. Pathania, Elements of Physical Chemistry, 3rd edn. Vishal Pub. CO., 2008.
6	B. R. Puri, L. R. Sharma, Kalia, Principles of Inorganic Chemistry, 31st edn. Milstone (2010).
7	.Manas Chanda, Atomic Structure and Molecular Spectroscopy.
8	Vogel's Text Book of Quantitative Chemical Analysis, J. Mendham, R. C. Denney, J.D. Barnes, M. Thomas, 6th edn. Pearson Education (2003).

SACRED HEART COLLEGE(AUTONOMOUS), THEVARA

DEPARTMENT OF CHEMISTRY

COURSE PLAN

ACADEMIC YEAR 2016-2017

PROGRAM ME	BSc BOTANY	SEMESTER	1
COURSE CODE AND TITLE	15U1PCHE1, GENERAL CHEMISTRY	CREDIT	2
HOURS/ SEM	36		
FACULTY NAME	Dr. Ramakrishnan S, Mr.Senju Devassykutty		

COURSE OUTCOMES (COs)	
1	Describe different models of atomic structure.
2	Define acids and bases and explain the concept of equilibrium.
3	Explain the fundamentals of nuclear chemistry and analytical chemistry.
4	Understand basics of thermodynamics

MODULE I : Atomic Structure : Dr. Ramakrishnan S				
Sl.No	Session	Topic	Method of Teaching	COs
1	1	Introduction: Atoms, Planck's quantum Theory, Photoelectric effect	Chalk & Board	1
2	2	Postulates of bohr's theory, Energy levels in atom	Chalk & Board	1
3	3	origin of hydrogen spectrum	Chalk & Board	1
4	4	Sommerfeld's extension of Bohr's Theory	Chalk & Board	1
5	5	Shortcomings of Bohr Theory	Chalk & Board	1
First Internal Examination				
6	6	Dual nature of matter and radiation. Derivation of de Broglie equation	Chalk & Board	1
7	7	Wave nature of electron and quantisation of angular momentum, Heisenberg's uncertainty principle	Chalk & Board	1
8	8	Concept of orbital, Quantum numbers, shapes of orbitals	Chalk & Board	1

9	9	Electronic configuration of atoms - Aufbau principle, Hund's rule of maximum multiplicity, Pauli's exclusion principle	Chalk & Board	1
MODULE II: Concept of Equilibrium: Dr. Ramakrishnan S				
1	1	Acids and bases	Chalk & Board	2
2	2	theories of acids and bases	Chalk & Board	2
3	3	ionic product of water, introductory idea of pH, pOH.	Chalk & Board	2
4	4	Strengths of acids and bases, K_a and K_b , pK_a and pK_b	Chalk & Board	2
5	5	buffer solution, Henderson equation	Chalk & Board	2
6	6	hydrolysis of salt, solubility	Chalk & Board	2
<i>Second Internal Examination</i>				
7	7	solubility product, application	Chalk & Board	2
8	8	Common ion effect, application.	Chalk & Board	2
MODULE III: Nuclear Chemistry : Mr.Senju Devassykutty				
1	1	Stability of Nucleus	Chalk & Board	3
2	2	natural radioactivity, induced radioactivity	Chalk & Board	3
3	3	fertile and fissile isotopes, units of radioactivity.	Chalk & Board	3
4	4	Nuclear Reactions: fission and fusion, chain reactions	Chalk & Board	3
5	5	disposal of nuclear wastes	Chalk & Board	3
6	6	applications of radioactivity	Chalk & Board	3
MODULE IV: Analytical Chemistry- Basic principles: Mr.Senju Devassykutty				
1	1	Concentration terms- molality, molarity, normality, weight percentage, ppm, and millimoles.	Chalk & Board	3
<i>First Internal Examination</i>				
2	2	Titrimetric method of analysis: General principle, types of titrations, requirements for titrimetric analysis.	Chalk & Board	3
3	3	Primary and secondary standards, criteria for primary standards, preparation of standard solutions, standardization of	Chalk & Board	3

		solutions		
MODULE V: Laws of Thermodynamics : Mr.Senju Devassykutty				
1	1	System and Surrounding	Chalk & Board	4
2	2	First law of Thermodynamics: Internal energy, Significance of internal energy change, enthalpy	Chalk & Board	4
3	3	Second law of Thermodynamics: free energy, Entropy and Spontaneity,	Chalk & Board	4
Second internal Examination				
4	4	Statement of second law based on entropy	Chalk & Board	4
5	5	Entropy change in Phase transitions	Chalk & Board	4
6	6	entropy of fusion, entropy of vaporization, entropy of sublimation	Chalk & Board	4
7	7	The concept of Gibbs's free energy- Physical significance of free energy, conditions for equilibrium & spontaneity based on ΔG values.	Chalk & Board	4
8	8	Effect of temperature on spontaneity of Reaction.	Chalk & Board	4
9	9	Third law of thermodynamics	Chalk & Board	4

TEXTBOOKS AND REFERENCES	
1	P. L. Soni, Inorganic Chemistry.
2	C. N. R. Rao, University General Chemistry, Macmillan.
3	R. A. Day Junior, A.L. Underwood, Quantitative Analysis, 5th edn. Prentice Hall of India Pvt. Ltd. New Delhi, 1988.
4	R. Gopalan, Analytical Chemistry, S. Chand and Co., New Delhi.
5	B. R. Puri, L. R. Sharma, M.S. Pathania, Elements of Physical Chemistry, 3rd edn. Vishal Pub. CO., 2008.
6	B. R. Puri, L. R. Sharma, Kalia, Principles of Inorganic Chemistry, 31st edn. Milstone (2010).
7	.Manas Chanda, Atomic Structure and Molecular Spectroscopy.
8	Vogel's Text Book of Quantitative Chemical Analysis, J. Mendham, R. C. Denney, J.D. Barnes, M. Thomas, 6th edn. Pearson Education (2003).

ASSIGNMENTS AND SEMINARS

Sl No	Module	Topic	Nature of Assignment	Alignment with POs, PSOs and COs
1	1	atomic structure	Numerical problems	CO1
2	2	concepts of chemical equilibrium	Numerical problems	CO2

SACRED HEART COLLEGE(AUTONOMOUS), THEVARA

DEPARTMENT OF CHEMISTRY

COURSE PLAN

ACADEMIC YEAR 2016-2017

PROGRAM ME	BSc Zoology	SEMESTER	1
COURSE CODE AND TITLE	15U1PCHE1, GENERAL CHEMISTRY	CREDIT	2
HOURS/ SEM	36		
FACULTY NAME	Dr. Ramakrishnan S, Mr.Senju Devassykutty		

COURSE OUTCOMES (COs)	
1	Describe different models of atomic structure.
2	Define acids and bases and explain the concept of equilibrium.
3	Explain the fundamentals of nuclear chemistry and analytical chemistry.
4	Understand basics of thermodynamics

Dr. Ramakrishnan S

MODULE I : Atomic Structure (9hrs)

Sl.No	Session	Topic	Method of Teaching	COs
1	1	Introduction: Atoms, Planck's quantum Theory, Photoelectric effect	Chalk & Board	1
2	2	Postulates of bohr's theory, Energy levels in atom	Chalk & Board	1
3	3	origin of hydrogen spectrum	Chalk & Board	1
4	4	Sommerfeld's extension of Bohr's Theory	Chalk & Board	1
5	5	Shortcomings of Bohr Theory	Chalk & Board	1
First Internal Examination				
6	6	Dual nature of matter and radiation. Derivation of de Broglie equation	Chalk & Board	1
7	7	Wave nature of electron and quantisation of angular momentum, Heisenberg's uncertainty principle	Chalk & Board	1
8	8	Concept of orbital, Quantum numbers,	Chalk & Board	1

		shapes of orbitals		
9	9	Electronic configuration of atoms - Aufbau principle, Hund's rule of maximum multiplicity, Pauli's exclusion principle	Chalk & Board	1
MODULE II: Concept of Equilibrium: (12 hrs)				
10	10	Acids and bases	Chalk & Board	2
11	11	theories of acids and bases	Chalk & Board	2
12	12	ionic product of water, introductory idea of pH, pOH.	Chalk & Board	2
13	13	Strengths of acids and bases, K_a and K_b , pK_a and pK_b	Chalk & Board	2
14	14	buffer solution, Henderson equation	Chalk & Board	2
15	15	hydrolysis of salt, solubility	Chalk & Board	2
Second Internal Examination				
16	16	solubility product, application	Chalk & Board	2
17	17	Common ion effect, application.	Chalk & Board	2
Mr.Senju Devassykutty				
MODULE III: Nuclear Chemistry : (6 hrs)				
1	1	Stability of Nucleus	Chalk & Board	3
2	2	natural radioactivity, induced radioactivity	Chalk & Board	3
3	3	fertile and fissile isotopes, units of radioactivity.	Chalk & Board	3
4	4	Nuclear Reactions: fission and fusion, chain reactions	Chalk & Board	3
5	5	disposal of nuclear wastes	Chalk & Board	3
6	6	applications of radioactivity	Chalk & Board	3
MODULE IV: Analytical Chemistry- Basic principles: (5hrs)				
7	7	Concentration terms- molality, molarity, normality, weight percentage, ppm, and millimoles.	Chalk & Board	3
First Internal Examination				
8	8	Titrimetric method of analysis: General principle, types of titrations, requirements for titrimetric analysis.	Chalk & Board	3
9	9	Primary and secondary standards, criteria	Chalk & Board	3

		for primary standards, preparation of standard solutions, standardization of solutions		
10	10	Errors	Chalk & Board	3
11	11	Minimization of errors	Chalk & Board	3
MODULE V: Laws of Thermodynamics : (8 hrs)				
12	12	System and Surrounding	Chalk & Board	4
13	13	First law of Thermodynamics: Internal energy, Significance of internal energy change, enthalpy	Chalk & Board	4
14	14	Second law of Thermodynamics: free energy, Entropy and Spontaneity,	Chalk & Board	4
Second internal Examination				
15	15	Statement of second law based on entropy	Chalk & Board	4
16	16	Entropy change in Phase transitions	Chalk & Board	4
17	17	entropy of fusion, entropy of vaporization, entropy of sublimation	Chalk & Board	4
18	18	The concept of Gibbs's free energy- Physical significance of free energy, conditions for equilibrium & spontaneity based on ΔG values.	Chalk & Board	4
19	19	Effect of temperature on spontaneity of Reaction. Third law of thermodynamics	Chalk & Board	4

TEXTBOOKS AND REFERENCES	
1	P. L. Soni, Inorganic Chemistry.
2	C. N. R. Rao, University General Chemistry, Macmillan.
3	R. A. Day Junior, A.L. Underwood, Quantitative Analysis, 5th edn. Prentice Hall of India Pvt. Ltd. New Delhi, 1988.
4	R. Gopalan, Analytical Chemistry, S. Chand and Co., New Delhi.
5	B. R. Puri, L. R. Sharma, M.S. Pathania, Elements of Physical Chemistry, 3rd edn. Vishal Pub. CO., 2008.
6	B. R. Puri, L. R. Sharma, Kalia, Principles of Inorganic Chemistry, 31st edn. Milstone (2010).
7	.Manas Chanda, Atomic Structure and Molecular Spectroscopy.
8	Vogel's Text Book of Quantitative Chemical Analysis, J. Mendham, R. C. Denney, J.D. Barnes, M. Thomas, 6th edn. Pearson Education (2003).

SACRED HEART COLLEGE(AUTONOMOUS), THEVARA

DEPARTMENT OF CHEMISTRY

COURSE PLAN

ACADEMIC YEAR 2016-2017

PROGRAMME	BSc PHYSICS	SEMESTER	2
COURSE CODE AND TITLE	15U2CPCHE2, BASIC ORGANIC CHEMISTRY	CREDIT	2
HOURS/ SEM	36		
FACULTY NAME	Dr. Ramakrishnan S, Mr.Senju Devassykutty		

COURSE OUTCOMES (COs)	
1	Understand the basics of organic chemistry.
2	Understand various purification techniques like solvent extraction, distillation and crystallization.
3	Explain the basics of organic reaction mechanism.
4	Discuss the classification and synthesis of polymers

Dr. Ramakrishnan S					
MODULE I: Purification of Organic Compounds:					
Sl.No	Session	Topic	Method of Teaching	Value Additions	CO
1	1	Purification techniques: Recrystallisation, sublimation.	Chalk & Board		
2	2	General principles of distillation, fractional distillation, distillation under reduced pressure	Chalk & Board		
3	3	Solvent extraction	Chalk & Board		

MODULE III: Mechanisms of Organic Reactions :					
4	4	Hybridization and shape of molecules - sp^3 , sp^2 and sp , (ethane, ethene, ethyne)	Chalk & Board		

5	5	Types of reagents - electrophiles, nucleophiles.	Chalk & Board		
6	6	Types of electron displacement in organic molecules	Chalk & Board		
7	7	Explanation of the strength of carboxylic and halogen substituted acids,	Chalk & Board		
8	8	base strength of primary, secondary and tertiary amines.	Chalk & Board		
9	9	Types of bond fission- homolytic and heterolytic fission.	Chalk & Board		
First Internal Examination					
10	10	Reactive intermediates- carbocations-Their formation and stability.	Chalk & Board		
11	11	Substitution reactions: Nucleophilic substitution of alkyl halides- S _N 1 and S _N 2 mechanisms. Factors affecting rate of Substitution reaction of alkyl halide. Nature of alkyl halide, Effect of solvent. Stereochemistry of S _N 1 and S _N 2 reactions.	Chalk & Board		
12	12	Electrophilic substitution in benzene- reaction mechanism. Halogenation, Nitration and Sulphonation	Chalk & Board		
13	13	Addition reactions: electrophilic addition. Addition of Bromine and Hydrogen halides to ethane, propene and ethyne-the Markwonikoff's rule, Peroxide effect.	Chalk & Board		
Second Internal Examination					
14	14	Reactive intermediates- free radicals. Their formation and stability.	Chalk & Board		
15	15	Substitution reactions: Nucleophilic substitution of alkyl halides- SN1 and SN2 mechanisms. Factors affecting rate of Substitution reaction of alkyl halide. Nature of alkyl halide, Effect of solvent. Stereochemistry of SN1 and SN2 reactions.	Chalk & Board		
16	16	Electrophilic substitution in benzene- reaction mechanism. Halogenation, Nitration and Sulphonation	Chalk & Board		
17	17	Addition reactions: electrophilic addition. Addition of Bromine and Hydrogen halides to ethane, propene and ethyne-the Markwonikoff's rule, Peroxide effect.	Chalk & Board		
18	18	Elimination reactions: E1 and E2 mechanisms. Saytzeff and Hofmann elimination	Chalk & Board		

Mr. Senju Devassykutty

MODULE II: Stereochemistry of Organic Compounds:

1	1	Geometrical isomerism- <i>cis</i> and <i>trans</i> configuration, 2-butene, maleic and fumaric acid,	Chalk & Board		
2	2	determination of configuration of cis-trans isomers, E and Z configuration.	Chalk & Board		
3	3	Optical isomerism- D and L configuration. Optical activity,	Chalk & Board		
4	4	Chirality, Stereogenic Centre	Chalk & Board		
5	5	Enantiomers and diastereomers	Chalk & Board		
6	6	optical isomerism in lactic acid and tartaric acid,	Chalk & Board		
First Internal examination					
7	7	Racemisation	Chalk & Board		
8	8	Conformation- Newman projection, Saw-horse projection	Chalk & Board		
9	9	Conformations of Ethane	Chalk & Board		
10	10	n - butane	Chalk & Board		
11	11	Cyclohexane.	Chalk & Board		

MODULE V: Natural and Synthetic Polymers:

12	12	Classification of polymers: Natural, synthetic; linear, cross-linked and network; plastics, elastomers, fibres; homopolymers and copolymers.	Chalk & Board		
13	13	Polymerization reactions, Addition Polymerization,	Chalk & Board		
Second Internal examination					
14	14	Condensation polymerization,	Chalk & Board		
15	15	polyethene, polypropylene, PVC, phenol-formaldehyde resins, polyamides (nylons) and polyester.	Chalk & Board		
16	16	Natural rubber: structure, vulcanization.	Chalk &		

			Board		
17	17	Synthetic rubbers- SBR, nitrile rubber, neoprene.	Chalk & Board		
18	18	Biodegradable polymers, environmental hazards caused by polymers, Health problem due to burning plastics.	Chalk & Board		

ASSIGNMENTS AND SEMINARS				
Sl No	Module	Topic	Nature of Assignment	Alignment with POs, PSOs and COs
1	1	Stereochemistry of Organic Compounds : some organic compounds	written	CO1
2	2	Natural rubber- structure, sources, processes, applications	written	CO4

TEXTBOOKS AND REFERENCES	
1	I. L. Finar, Organic Chemistry, Vol. I, 6th edn. Pearson.
2	S. M. Mukherji, S. P Singh, R. P Kapoor, Organic Chemistry, Vol.1, New Age International (P) Ltd, 2006
3	P.S Kalsi, Stereochemistry Conformation and Mechanism, New Age International Publishers, 2004
4	Peter Sykes, A Guide Book to Mechanism in Organic Chemistry, 6th edn. Orient Longman, 1988
5	S. M. Mukherji, S.P Singh, Reaction Mechanism in Organic Chemistry, Macmillan, 3rd Edn., 2003
6	V. R. Gowariker, Polymer Science, Wiley Eastern.
7	K.S Tewari, N K Vishnoi, Text book of Organic Chemistry, Vikas Publishing House Pvt. Ltd.2007.

SACRED HEART COLLEGE(AUTONOMOUS), THEVARA

DEPARTMENT OF CHEMISTRY

COURSE PLAN

ACADEMIC YEAR 2016-2017

PROGRAMME	BSc BOTANY	SEMESTER	2
COURSE CODE AND TITLE	15U2CPCHE2, BASIC ORGANIC CHEMISTRY	CREDIT	2
HOURS/ SEM	36		
FACULTY NAME	Dr. Ramakrishnan S, Mr.Senju Devassykutty		

COURSE OUTCOMES (COs)	
1	Understand the basics of organic chemistry.
2	Understand various purification techniques like solvent extraction, distillation and crystallization.
3	Explain the basics of organic reaction mechanism.
4	Discuss the classification and synthesis of polymers

MODULE I: Purification of Organic Compounds: Dr. Ramakrishnan S					
Sl.No	Session	Topic	Method of Teaching	Value Additions	CO
1	1	Purification techniques: Recrystallisation, sublimation.	Chalk & Board		
2	2	General principles of distillation, fractional distillation, distillation under reduced pressure	Chalk & Board		
3	3	Solvent extraction	Chalk & Board		

MODULE III: Mechanisms of Organic Reactions : Dr. Ramakrishnan S.					
4	4	Hybridization and shape of molecules - sp^3 , sp^2 and sp , (ethane, ethene, ethyne)			

5	5	Types of reagents - electrophiles, nucleophiles.	Chalk & Board		
6	6	Types of electron displacement in organic molecules	Chalk & Board		
7	7	Explanation of the strength of carboxylic and halogen substituted acids,	Chalk & Board		
8	8	base strength of primary, secondary and tertiary amines.	Chalk & Board		
9	9	Types of bond fission- homolytic and heterolytic fission.	Chalk & Board		
First Internal Examination					
10		Reactive intermediates- carbocations-Their formation and stability.	Chalk & Board		
11		Substitution reactions: Nucleophilic substitution of alkyl halides- S _N 1 and S _N 2 mechanisms. Factors affecting rate of Substitution reaction of alkyl halide. Nature of alkyl halide, Effect of solvent. Stereochemistry of S _N 1 and S _N 2 reactions.	Chalk & Board		
12		Electrophilic substitution in benzene- reaction mechanism. Halogenation, Nitration and Sulphonation	Chalk & Board		
13		Addition reactions: electrophilic addition. Addition of Bromine and Hydrogen halides to ethane, propene and ethyne-the Markwonikoff's rule, Peroxide effect.	Chalk & Board		
Second Internal Examination					
14		Reactive intermediates- free radicals. Their formation and stability.	Chalk & Board		
15		Substitution reactions: Nucleophilic substitution of alkyl halides- SN1 and SN2 mechanisms. Factors affecting rate of Substitution reaction of alkyl halide. Nature of alkyl halide, Effect of solvent. Stereochemistry of SN1 and SN2 reactions.	Chalk & Board		
16		Electrophilic substitution in benzene- reaction mechanism. Halogenation, Nitration and Sulphonation	Chalk & Board		
17		Addition reactions: electrophilic addition. Addition of Bromine and Hydrogen halides to ethane, propene and ethyne-the Markwonikoff's rule, Peroxide effect.	Chalk & Board		
18		Elimination reactions: E1 and E2 mechanisms. Saytzeff and Hofmann elimination	Chalk & Board		

MODULE II: Stereochemistry of Organic Compounds: Mr. Senju Devassykutty					
1	1	Geometrical isomerism- <i>cis</i> and <i>trans</i> configuration, 2-butene, maleic and fumaric acid,	Chalk & Board		
2	2	determination of configuration of cis-trans isomers, E and Z configuration.	Chalk & Board		
3	3	Optical isomerism- D and L configuration. Optical activity,	Chalk & Board		
4	4	Chirality, Stereogenic Centre	Chalk & Board		
5	5	Enantiomers and diastereomers	Chalk & Board		
6	6	optical isomerism in lactic acid and tartaric acid,	Chalk & Board		
First Internal examination					
7	7	Racemisation	Chalk & Board		
8	8	Conformation- Newman projection, Saw-horse projection	Chalk & Board		
9	9	Conformations of Ethane	Chalk & Board		
10	10	n - butane	Chalk & Board		
11	11	Cyclohexane.	Chalk & Board		

MODULE V: Natural and Synthetic Polymers: Mr. Senju Devassykutty					
12	12	Classification of polymers: Natural, synthetic; linear, cross-linked and network; plastics, elastomers, fibres; homopolymers and copolymers.	Chalk & Board		
13	13	Polymerization reactions, Addition Polymerization,	Chalk & Board		
Second Internal examination					
14	14	Condensation polymerization,	Chalk & Board		
15	15	polyethene, polypropylene, PVC, phenol-formaldehyde resins, polyamides (nylons) and polyester.	Chalk & Board		
16	16	Natural rubber: structure, vulcanization.	Chalk & Board		
17	17	Synthetic rubbers- SBR, nitrile rubber,	Chalk &		

		neoprene.	Board		
18	18	Biodegradable polymers, environmental hazards caused by polymers, Health problem due to burning plastics.	Chalk & Board		

ASSIGNMENTS AND SEMINARS				
Sl No	Module	Topic	Nature of Assignment	Alignment with POs, PSOs and COs
1	1	Stereochemistry of Organic Compounds : some organic compounds	written	CO1
2	2	Natural rubber- structure, sources, processes, applications	written	CO4

TEXTBOOKS AND REFERENCES	
1	I. L. Finar, Organic Chemistry, Vol. I, 6th edn. Pearson.
2	S. M. Mukherji, S. P Singh, R. P Kapoor, Organic Chemistry, Vol.1, New Age International (P) Ltd, 2006
3	P.S Kalsi, Stereochemistry Conformation and Mechanism, New Age International Publishers, 2004
4	Peter Sykes, A Guide Book to Mechanism in Organic Chemistry, 6th edn. Orient Longman, 1988
5	S. M. Mukherji, S.P Singh, Reaction Mechanism in Organic Chemistry, Macmillan, 3rd Edn., 2003
6	V. R. Gowariker, Polymer Science, Wiley Eastern.
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SACRED HEART COLLEGE(AUTONOMOUS), THEVARA

DEPARTMENT OF CHEMISTRY

COURSE PLAN

ACADEMIC YEAR 2016-2017

PROGRAMME	BSc Zoology	SEMESTER	2
COURSE CODE AND TITLE	15U2CPCHE2, BASIC ORGANIC CHEMISTRY	CREDIT	2
HOURS/ SEM	36		
FACULTY NAME	Dr. Ramakrishnan S, Mr.Senju Devassykutty		

COURSE OUTCOMES (COs)	
1	Understand the basics of organic chemistry.
2	Understand various purification techniques like solvent extraction, distillation and crystallization.
3	Explain the basics of organic reaction mechanism.
4	Discuss the classification and synthesis of polymers

Dr. Ramakrishnan S					
MODULE I: Purification of Organic Compounds:					
Sl.No	Session	Topic	Method of Teaching	Value Additions	CO
1	1	Purification techniques: Recrystallisation, sublimation.	Chalk & Board		
2	2	General principles of distillation, fractional distillation, distillation under reduced pressure	Chalk & Board		
3	3	Solvent extraction	Chalk & Board		

MODULE III: Mechanisms of Organic Reactions :					
4	4	Hybridization and shape of molecules - sp^3 , sp^2 and sp , (ethane, ethene, ethyne)	Chalk & Board		

5	5	Types of reagents - electrophiles, nucleophiles.	Chalk & Board		
6	6	Types of electron displacement in organic molecules	Chalk & Board		
7	7	Explanation of the strength of carboxylic and halogen substituted acids,	Chalk & Board		
8	8	base strength of primary, secondary and tertiary amines.	Chalk & Board		
9	9	Types of bond fission- homolytic and heterolytic fission.	Chalk & Board		
First Internal Examination					
10	10	Reactive intermediates- carbocations-Their formation and stability.	Chalk & Board		
11	11	Substitution reactions: Nucleophilic substitution of alkyl halides- S _N 1 and S _N 2 mechanisms. Factors affecting rate of Substitution reaction of alkyl halide. Nature of alkyl halide, Effect of solvent. Stereochemistry of S _N 1 and S _N 2 reactions.	Chalk & Board		
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16	16	Electrophilic substitution in benzene- reaction mechanism. Halogenation, Nitration and Sulphonation	Chalk & Board		
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18	18	Elimination reactions: E1 and E2 mechanisms. Saytzeff and Hofmann elimination	Chalk & Board		

Mr. Senju Devassykutty

MODULE II: Stereochemistry of Organic Compounds:

1	1	Geometrical isomerism- <i>cis</i> and <i>trans</i> configuration, 2-butene, maleic and fumaric acid,	Chalk & Board		
2	2	determination of configuration of cis-trans isomers, E and Z configuration.	Chalk & Board		
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6	6	optical isomerism in lactic acid and tartaric acid,	Chalk & Board		
First Internal examination					
7	7	Racemisation	Chalk & Board		
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9	9	Conformations of Ethane	Chalk & Board		
10	10	n - butane	Chalk & Board		
11	11	Cyclohexane.	Chalk & Board		

MODULE V: Natural and Synthetic Polymers:

12	12	Classification of polymers: Natural, synthetic; linear, cross-linked and network; plastics, elastomers, fibres; homopolymers and copolymers.	Chalk & Board		
13	13	Polymerization reactions, Addition Polymerization,	Chalk & Board		
Second Internal examination					
14	14	Condensation polymerization,	Chalk & Board		
15	15	polyethene, polypropylene, PVC, phenol-formaldehyde resins, polyamides (nylons) and polyester.	Chalk & Board		
16	16	Natural rubber: structure, vulcanization.	Chalk &		

			Board		
17	17	Synthetic rubbers- SBR, nitrile rubber, neoprene.	Chalk & Board		
18	18	Biodegradable polymers, environmental hazards caused by polymers, Health problem due to burning plastics.	Chalk & Board		

ASSIGNMENTS AND SEMINARS				
Sl No	Module	Topic	Nature of Assignment	Alignment with POs, PSOs and COs
1	1	Stereochemistry of Organic Compounds : some organic compounds	written	CO1
2	2	Natural rubber- structure, sources, processes, applications	written	CO4

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1	I. L. Finar, Organic Chemistry, Vol. I, 6th edn. Pearson.
2	S. M. Mukherji, S. P Singh, R. P Kapoor, Organic Chemistry, Vol.1, New Age International (P) Ltd, 2006
3	P.S Kalsi, Stereochemistry Conformation and Mechanism, New Age International Publishers, 2004
4	Peter Sykes, A Guide Book to Mechanism in Organic Chemistry, 6th edn. Orient Longman, 1988
5	S. M. Mukherji, S.P Singh, Reaction Mechanism in Organic Chemistry, Macmillan, 3rd Edn., 2003
6	V. R. Gowariker, Polymer Science, Wiley Eastern.
7	K.S Tewari, N K Vishnoi, Text book of Organic Chemistry, Vikas Publishing House Pvt. Ltd.2007.

COURSE PLAN					
ACADEMIC YEAR 2016-17					
PROGRAMME	:	<i>B.Sc. Physics</i>	LECTURE HOURS	:	54
SEMESTER	:	3	CREDITS	:	3
SUBJECT TITLE	:	ADVANCED PHYSICAL CHEMISTRY – I	SUBJECT CODE	:	15U3CPCHE3. 1
COURSE TEACHERS	:	<i>Dr. MIDHUN DOMINIC C.D, Dr. THOMMACHAN XAVIER, Dr. K.B.JOSE</i>			
COURSE OUTCOMES (COs)	:	CO1 - To have idea about Naochemistry and nanotechnology CO2 - Determination of symmetry of molecules CO3 - Study of phase equilibria of different systems CO4 - An idea of solid state ad crystal systems CO5 - An idea on properties of liquid state CO6 - To study the basic concepts of surface chemistry CO7 - To have an idea about colloids			

Dr. MIDHUN DOMINIC C D				
Unit	No. of Sessions	Session Topic and Discussion Theme	Mode of teaching	COs
Nanomaterials and Nanotechnology (8 hrs)	1	Introduction to nanoscience-Moore's law	Lecture	CO1 CO1 CO1 CO1 CO5 CO1
	2	properties of nanomaterials	Lecture	
	3	Synthesis of nanomaterials-reduction method-precipitation method-sol gel method	Lecture	
	4	Green synthesis of nanosilver and nanogold-	Lecture	
Liquid State(5 hrs)	First Internal examination			
	5	Properties and applications of nanomaterials	Lecture	CO5 CO5
	6	nanocomposites-nanomedicine.	Lecture	
	7	Intermolecular forces liquids compared with gases and solids	Lecture	
8				

	8	viscosity, surface tension	Lecture	
	9	Liquid crystals – the intermediate phase between solid and normal liquid phases,	Lecture	
	Second Internal examination			
	10	thermographic behavior, classification	Lecture	CO5
	11	structure of nematic and cholesteric phases.	Lecture	CO5 CO6
	12	Adsorption – types of adsorption of gases by solids	Lecture	CO6 CO6
	13	Factors influencing adsorption	Lecture	
	14	Freundlich adsorption isotherm – Langmuir adsorption isotherm	Lecture	CO7 CO7
	15	Colloids: preparation, properties – optical and electrical	Lecture	CO7
	16	Electric double layer, coagulation, electrophoresis, electro osmosis	Lecture	
	17	Surfactants, micelle, applications of colloids	Seminar	
	References: <ol style="list-style-type: none"> 1. B. R. Puri, L. R. Sharma, M. S. Pathania, Elements of Physical Chemistry, 40th edn. Vishal Pub. Co. Jalandhar (2003) 2. Ashcroft / Mermin, Solid State Physics, Thomson Publishers 3. J. Tareen and T. Kutty, A basic course in Crystallography, University Press. 			
<i>Dr. THOMMACHAN XAVIER</i>				
Symmetry and Molecular Structure (9 hrs)	1	Symmetry elements and symmetry operation	Lecture	CO2 CO2
	2	Centre of symmetry, plane of symmetry,	Lecture	CO2
	3	proper and improper axes of symmetry,	Lecture	CO2
	4	identity, molecular point groups,	Lecture	CO2
	5	Schoeniflies symbol and determination of point groups of simple molecule- H ₂ O	Lecture	CO2
	6	point groups of simple molecule NH ₃ , BF ₃ ,	Lecture	
	First Internal examination			
	7	point groups of simple molecule CO, HCl,	Lecture	CO2 CO2
	8	point groups of simple molecule C ₂ H ₂ , CO ₂ ,	Lecture	
	9	point groups of simple molecule Benzene, NO ₃ ⁻ , PCl ₅ .	Lecture	
	10	Revision	Discussion	
Phase Equilib	11	The phase rule, definition,	Lecture	CO3
	12	equilibrium between phases, one component system – water system	Lecture	CO3 CO3

rium (9 hrs)	13	Sulphur system.	Lecture	CO3
	14	Distribution law, partition coefficient,	Lecture	
	Second Internal examination			
	15	applications- Study of association or dissociation	Lecture	CO3
	16	Principle of extraction. Distribution indicators.	Lecture	CO3 CO3
	17	Revision	Discussion	

References:

1. B. R. Puri, L. R. Sharma, M. S. Pathania, Elements of Physical Chemistry, 40th edn. Vishal Pub. Co. Jalandhar (2003)
2. Ashcroft / Mermin, Solid State Physics, Thomson Publishers
3. J. Tareen and T. Kutty, A basic course in Crystallography, University Press.

Dr. K B JOSE

	1	Classification: amorphous, crystalline – differences.	Lecture	CO4 CO4 CO4 CO4
	2	Lattice ,lattice energy (general idea)	Lecture	
	3	unit cell, examples of simple cubic	Lecture	
	4	bcc and fcc lattices	Lecture	
	5	Calculation of number of molecules in a unit cell.	Lecture	CO4 CO4
	6	Weiss and Miller indices, crystal systems	Lecture	
	First Internal examination			
	7	Bravais lattices, X-ray diffraction – Bragg’s equation,	Lecture	CO4 CO4 CO4 CO4 CO4 CO4 CO4
	8	structure determination of NaCl by X-ray diffraction.	Lecture	
	9	Theories of Solid: metallic bond	Lecture	
	10	band theory, conductors,	Lecture	
	11	semiconductors and insulators,	Lecture	
	12	mention of super conductors.	Lecture	
	13	Defects in solids-stoichiometric	Lecture	
14	Non-stoichiometric defects and	Lecture		

		consequences		
	<i>Second Internal examination</i>			
	15	Magnetic Properties: classification	Lecture	CO4
	16	Diamagnetic, paramagnetic,	Lecture	CO4
	17	Antiferromagnetic, ferro and ferrimagnetic,	Lecture	CO4
	18	permanent and temporary magnets.	Lecture	CO4
	19	Revision	Discussion	

Head of the Department

Thevara

Date 2/11/2018

COURSE PLAN					
ACADEMIC YEAR 2016-17					
PROGRAMME	:	<i>B.Sc. Botany</i>	LECTURE HOURS	:	<i>54</i>
SEMESTER	:	<i>3</i>	CREDITS	:	<i>3</i>
SUBJECT TITLE	:	INORGANIC AND BIO-INORGANIC CHEMISTRY	SUBJECT CODE	:	<i>15U3PCHE3.2</i>
COURSE TEACHERS	:	<i>Dr. Ramakrishnan S., Dr. Franklin John, Dr. Midhun Dominic</i>			
COURSE OUTCOMES (COs)	:	CO1 - <i>To have idea about Chemistry in Agriculture</i> CO2 - <i>Understand basic chemistry involved in biological processes</i> CO3 - <i>To have basic knowledge on some important enzymes and nucleic acids</i> CO4 - <i>To have idea about Chemistry in medicine</i>			

Dr. Ramakrishnan S				
Unit	No. of Sessions	Session Topic and Discussion Theme	Mode of teaching	COs
Chemistry and Agriculture (12hrs)	1	Mineral nutrients-Macro nutrients- Primary and Secondary macro nutrients.	Lecture	CO1
	2	Micronutrients. Their role in plant growth.	Lecture	
	3	Pesticides-classifications simple examples,	Lecture	
	4	NPK value superphosphates, triple super phosphate	Lecture	
	5	uses of mixed fertilizers,	Lecture	
	6	Bio-fertilizers.	Lecture	
	<i>First Internal examination</i>			
7	Plant growth hormones.	Lecture		

	8	bio pesticides	Lecture	
	9	Insecticides – stomach poisons,	Lecture	
	10	contact insecticides, Examples	Lecture	
	11	fumigants. Examples	Lecture	
	12	Method of preparation of DDT, BHC, pyrethrin.	Lecture	
<p>Reference:</p> <p>1. R. Puri, L. R. Sharma, K. C. Kalia, Principles of Inorganic Chemistry, 31st Milestone Publishers, New Delhi, 2010.</p> <p>2. http://en.wikipedia.org/wiki/Plant_nutrition.</p>				
Chemistry in medicine (6 hrs)	13	antibiotics	Lecture	CO4
	14	sulpha drugs	Lecture	
	15	antipyretics	Lecture	
	<i>Second Internal examination</i>			
	16	analgesics	Lecture	
	17	tranquillizers, and antidepressants	Lecture	
	18	Drug addiction abuse and prevention.	Lecture	
<p>Reference:</p> <p>1. I. L. Finar, Organic Chemistry, Vol. 1 & 2, 6th Edition, Pearson.</p> <p>2. C. N. R. Rao, University General Chemistry, Macmillan.</p> <p>3. G. T. Austin, Shreve's Chemical process Industries, 5th edition, McGraw Hill, 1984.</p> <p>4. Rastogi, Biochemistry, Tata McGraw. Hill Publication, 1996.</p>				
Dr. Franklin John				
Oxygen carriers in	1	Oxygen transport in biological system- Hemoglobin and myoglobin	Lecture	CO2
	2	Structure and function of Hemoglobin and myoglobin	Lecture	
	3	Oxygen transport mechanism	Lecture	
	4	cooperativity of hemoglobin	Lecture	
	5	Perutz mechanism	Lecture	
	6	Bohr effect.	Lecture	
	<i>First Internal examination</i>			
	7	Hemocyanin, Hemerythrin	Lecture	
8	Photosynthetic pigments, Chlorophyll,	Lecture		

biological system hr 9)		Structure			
	9	Different types of chlorophyll			
	10	Photosystem-I, photosystem-II			
	11	Z- Scheme	Lecture		
	12	photophosphorylation	Discussion		
Chemistry in medicine (6 hrs)	13	Essential and trace elements in biological system.	Lecture	CO4	
	14	Metal deficiency and disease- Iron, Zinc Copper	Lecture		
	15	Metal toxicity- Toxicity due to Copper overload	Lecture		
	16	Metal toxicity- Toxicity due Plutonium	Lecture		
	Second Internal examination				
	17	Metal toxicity- Toxicity due mercury	Lecture		
	18	Metal toxicity- Toxicity due Iron	Discussion		
Reference					
J. D. Lee, Concise Inorganic Chemistry 5th edn., Wiley India Pvt. Ltd.2008.					
2. R. Puri, L. R. Sharma, K. C. Kalia, Principles of Inorganic Chemistry, 31 st Milestone Publishers, New Delhi, 2010.					
1. G.L. Meissler, D.A Tarr, Inorganic Chemistry, 3 rd Edn. Pearson Education, 2004.					
2. J. E. Huheey, E. A. Keiter, R. L. Keiter, O. K. Medhi, Inorganic Chemistry, Pearson 2006.					
Dr. Midhun Dominic					
Chemistry in medicine (3 hrs)	1	Treatment for excess iron, copper, mercury and plutonium.	Lecture	CO4	
	2	Anti-cancer drugs- Cisplatin.	Lecture		
	3	Anti-Arthritis drugs- Gold compounds	Lecture		
Enzymes and	4	Thermodynamics of Living cell- Exergonic and endergonic reactions	Lecture	CO3	
	5	coupled reactions, biological oxidation reactions	Lecture		
	Second Internal examination				
	6	Enzymes: classification, Metalloenzymes,	Lecture		

Nucleic acids (15 hrs)	7	prosthetic group, coenzyme, cofactors	Lecture
	8	characteristics of enzyme action	Lecture
	9	mechanism of enzyme action	Lecture
	10	Structure and function of some important enzymes	Lecture
	11	Structure and function of some important enzymes	Lecture
	12	Structure and function of some important enzymes	Lecture
	13	Structure and function of some important enzymes	Lecture
	14	Energy rich molecules: elementary structure of ATP and ADP.	Lecture
	15	Na ⁺ /K ⁺ ATPase-Sodium Potassium pump	Lecture
	16	Nucleic acids - Chemical composition,	Lecture
	<i>Second Internal examination</i>		
	17	Structure of DNA & RNA.	Lecture
	18	Biological Functions.	Lecture
	Reference 1. I. L. Finar, Organic Chemistry, Vol. 1 & 2, 6th Edition, Pearson. 2. C. N. R. Rao, University General Chemistry, Macmillan. 3. G. T. Austin, Shreve's Chemical process Industries, 5th edition, McGraw Hill, 1984. 4. Rastogi, Biochemistry, Tata McGraw. Hill Publication, 1996.		

Head of the Department

Thevara

Date

COURSE PLAN					
ACADEMIC YEAR 2016-17					
PROGRAMME	:	<i>B.Sc. Zoology</i>	LECTURE HOURS	:	<i>54</i>
SEMESTER	:	<i>3</i>	CREDITS	:	<i>3</i>
SUBJECT TITLE	:	INORGANIC AND BIO-INORGANIC CHEMISTRY	SUBJECT CODE	:	<i>15U3PCHE3.2</i>
COURSE TEACHERS	:	<i>Dr. Ramakrishnan S., Dr. Franklin John, Dr. Midhun Dominic</i>			
COURSE OUTCOMES (COs)	:	<p>CO1 - <i>To have idea about Chemistry in Agriculture</i></p> <p>CO2 - <i>Understand basic chemistry involved in biological processes</i></p> <p>CO3 - <i>To have basic knowledge on some important enzymes and nucleic acids</i></p> <p>CO4 - <i>To have idea about Chemistry in medicine</i></p>			

Dr. Ramakrishnan S				
Unit	No. of Sessions	Session Topic and Discussion Theme	Mode of teaching	COs
Chemistry and Agriculture (12hrs)	1	Mineral nutrients-Macro nutrients- Primary and Secondary macro nutrients.	Lecture	CO1
	2	Micronutrients. Their role in plant growth.	Lecture	
	3	Pesticides-classifications simple examples,	Lecture	
	4	NPK value superphosphates, triple super phosphate	Lecture	
	5	uses of mixed fertilizers,	Lecture	
	6	Bio-fertilizers.	Lecture	
	<i>First Internal examination</i>			
7	Plant growth hormones.	Lecture		

	8	bio pesticides	Lecture	
	9	Insecticides – stomach poisons,	Lecture	
	10	contact insecticides, Examples	Lecture	
	11	fumigants. Examples	Lecture	
	12	Method of preparation of DDT, BHC, pyrethrin.	Lecture	
<p>Reference:</p> <p>1. R. Puri, L. R. Sharma, K. C. Kalia, Principles of Inorganic Chemistry, 31st Milestone Publishers, New Delhi, 2010.</p> <p>2. http://en.wikipedia.org/wiki/Plant_nutrition.</p>				
Chemistry in medicine (6 hrs)	13	antibiotics	Lecture	CO4
	14	sulpha drugs	Lecture	
	15	antipyretics	Lecture	
	<i>Second Internal examination</i>			
	16	analgesics	Lecture	
	17	tranquillizers, and antidepressants	Lecture	
	18	Drug addiction abuse and prevention.	Lecture	
<p>Reference:</p> <p>1. I. L. Finar, Organic Chemistry, Vol. 1 & 2, 6th Edition, Pearson.</p> <p>2. C. N. R. Rao, University General Chemistry, Macmillan.</p> <p>3. G. T. Austin, Shreve's Chemical process Industries, 5th edition, McGraw Hill, 1984.</p> <p>4. Rastogi, Biochemistry, Tata McGraw. Hill Publication, 1996.</p>				
Dr. Franklin John				
Oxygen carriers in	1	Oxygen transport in biological system- Hemoglobin and myoglobin	Lecture	CO2
	2	Structure and function of Hemoglobin and myoglobin	Lecture	
	3	Oxygen transport mechanism	Lecture	
	4	cooperativity of hemoglobin	Lecture	
	5	Perutz mechanism	Lecture	
	6	Bohr effect.	Lecture	
	<i>First Internal examination</i>			
	7	Hemocyanin, Hemerythrin	Lecture	
8	Photosynthetic pigments, Chlorophyll,	Lecture		

biological system hr 9)		Structure			
	9	Different types of chlorophyll			
	10	Photosystem-I, photosystem-II			
	11	Z- Scheme	Lecture		
	12	photophosphorylation	Discussion		
Chemistry in medicine (6 hrs)	13	Essential and trace elements in biological system.	Lecture	CO4	
	14	Metal deficiency and disease- Iron, Zinc Copper	Lecture		
	15	Metal toxicity- Toxicity due to Copper overload	Lecture		
	16	Metal toxicity- Toxicity due Plutonium	Lecture		
	Second Internal examination				
	17	Metal toxicity- Toxicity due mercury	Lecture		
	18	Metal toxicity- Toxicity due Iron	Discussion		
Reference					
J. D. Lee, Concise Inorganic Chemistry 5th edn., Wiley India Pvt. Ltd.2008.					
2. R. Puri, L. R. Sharma, K. C. Kalia, Principles of Inorganic Chemistry, 31 st Milestone Publishers, New Delhi, 2010.					
1. G.L. Meissler, D.A Tarr, Inorganic Chemistry, 3 rd Edn. Pearson Education, 2004.					
2. J. E. Huheey, E. A. Keiter, R. L. Keiter, O. K. Medhi, Inorganic Chemistry, Pearson 2006.					
Dr. Midhun Dominic					
Chemistry in medicine (3 hrs)	1	Treatment for excess iron, copper, mercury and plutonium.	Lecture	CO4	
	2	Anti-cancer drugs- Cisplatin.	Lecture		
	3	Anti-Arthritis drugs- Gold compounds	Lecture		
Enzymes and	4	Thermodynamics of Living cell- Exergonic and endergonic reactions	Lecture	CO3	
	5	coupled reactions, biological oxidation reactions	Lecture		
	Second Internal examination				
	6	Enzymes: classification, Metalloenzymes,	Lecture		

Nucleic acids (15 hrs)	7	prosthetic group, coenzyme, cofactors	Lecture
	8	characteristics of enzyme action	Lecture
	9	mechanism of enzyme action	Lecture
	10	Structure and function of some important enzymes	Lecture
	11	Structure and function of some important enzymes	Lecture
	12	Structure and function of some important enzymes	Lecture
	13	Structure and function of some important enzymes	Lecture
	14	Energy rich molecules: elementary structure of ATP and ADP.	Lecture
	15	Na ⁺ /K ⁺ ATPase-Sodium Potassium pump	Lecture
	16	Nucleic acids - Chemical composition,	Lecture
	<i>Second Internal examination</i>		
	17	Structure of DNA & RNA.	Lecture
	18	Biological Functions.	Lecture
	Reference 1. I. L. Finar, Organic Chemistry, Vol. 1 & 2, 6th Edition, Pearson. 2. C. N. R. Rao, University General Chemistry, Macmillan. 3. G. T. Austin, Shreve's Chemical process Industries, 5th edition, McGraw Hill, 1984. 4. Rastogi, Biochemistry, Tata McGraw. Hill Publication, 1996.		

Head of the Department

Thevara

Date

COURSE PLAN					
ACADEMIC YEAR 2016-17					
PROGRAMME	:	<i>B.Sc. Physics</i>	LECTURE HOURS	:	54
SEMESTER	:	4	CREDITS	:	3
SUBJECT TITLE	:	ADVANCED PHYSICAL CHEMISTRY – II	SUBJECT CODE	:	15U4CPCHE4.1
COURSE TEACHERS	:	<i>Dr. MIDHUN DOMINIC C.D, Dr.THOMMACHAN XAVIERTX, Dr. K B JOSE</i>			
COURSE OUTCOMES (COs)	:	CO1 -To understand the basics of spectroscopy CO2 - To know the fundamental factors of chemical kinetics CO3 - To have an idea about photochemistry CO4 - basic understanding of electrochemistry CO5 - To understand the fundamental aspects of redox reactions			

Dr. MIDHUN DOMINIC C D				
Unit	No. of Sessions	Session Topic and Discussion Theme	Mode of teaching	COs
Introduction to Spectroscopy (12 hrs)	1	Interaction of electromagnetic radiation with matter, electromagnetic spectrum	Lecture	CO1 CO1
	2	Quantization of energy, electronic, vibrational and rotational energy levels	Lecture	CO1
	3	Boltzmann distribution of energy (formula only), population of levels	Lecture	CO1 CO1
	4	UV- Visible Spectroscopy: Beer Lambert's law	Lecture	CO1 CO1 CO1
	5	molar extinction coefficient and its importance	Lecture	CO1 CO1 CO1
	6	UV spectrum, max, chromophore, auxochrome	Lecture	CO1
	7	Red shift, blue shift, types of transition	Lecture	CO1
	8	Infra-red spectroscopy: vibrational degrees of freedom	Lecture	CO1 CO1 CO2
	9	Types of vibrations – symmetric and asymmetric stretching and bending	Lecture	
<i>First Internal examination</i>				

Chemical Kinetics (8hrs)

	10	Concept of group frequencies-frequencies of common functional groups in organic compounds.	Lecture	CO2
	11	Rotational Spectroscopy: diatomic molecules	Lecture	CO2
	12	Determination of bond length	Lecture	
	13	Rate of reaction, rate law	Lecture	CO2
	14	Order of reaction, molecularity of reaction	Lecture	CO2
	15	Integrated rate expression for first order reaction	Lecture	CO2
	16	Half life, determination of order of reactions	Lecture	
	<i>Second Internal examination</i>			
	17	Influence of temperature on reaction rate – Arrhenius equation	Lecture	
	18	Concept of activation energy	Lecture	
	19	Importance of activated complex	Lecture	
20	Catalysis, examples	Lecture		

Dr. THOMMACHAN XAVIER

Photochemistry (5 hrs)	1	Laws of Photochemistry	Lecture	CO3
	2	Photochemical process – primary and secondary, quantum yield	Lecture	CO3
	3	Basic Concepts of Photosensitized reactions	Lecture	CO3
	4	Flash photolysis and chemiluminescence	Lecture	CO3
	5	Frank-Condon principle – fluorescence and phosphorescence	Lecture	
Electrochemistry (12 hrs)	6	Conductance of electrolytic solution, electrolytic conductivity (K)		CO4
	7	Molar conductivity of solutions of electrolytes	Lecture	CO4
	<i>First Internal examination</i>			
	8	Variation of conductivity and molar conductivity with concentration	Lecture	CO4
	9	Kohlrausch's law – application	Lecture	CO4
	10	Faraday's laws of electrolysis	Lecture	CO4
	11	Electrochemical equivalent and chemical equivalent	Lecture	CO4
	12	Transport number-determination by Hittorf's method	Lecture	CO4
13	Applications of conductance	Lecture	CO4	

		measurements		CO4
	<i>Second Internal examination</i>			
	14	Kw, Ksp, Conductometric titrations	Lecture	CO4
	15	Strong and weak electrolytes.	Lecture	
	16	Ostwald's dilution law	Lecture	
	17	Hydrolysis of salts	Lecture	
Dr. K B JOSE				
Electromotive force (11 hrs)	1	Galvanic cells, characteristics of reversible cells	Lecture	CO4 CO4
	2	Reversible electrodes – different types	Lecture	CO4
	3	Electrode potential – effect of electrolyte concentration on electrode potential and emf (Nernst equation)	Lecture	
	4	Electrochemical series, representation of cell	Discussion	
	<i>First Internal examination</i>			
	5	EMF of cell	Lecture	CO4 CO4
	6	EMF and equilibrium constant of cell reaction	Lecture	CO4 CO4
	7	Concentration cells	Lecture	
	8	General discussion of electrode – concentration cell and electrolyte concentration cells	Lecture	CO4 CO4
	9	Liquid junction potential, fuel cells – the hydrogen – oxygen fuel cell.	Lecture	CO4
	10	Application of emf measurement – determination of pH using hydrogen electrode, quinhydrone electrode, glass electrode	Lecture	
Redox reactions	12	Oxidation Reduction reactions: explanation with examples	Lecture	CO5 CO5
	13	oxidation states	Lecture	
	<i>Second Internal examination</i>			CO5
	14	Rules to assign oxidation states in	Lecture	

CO5

CO5

(6 hrs)		polyatomic molecules		
	15	Determination of oxidation states	Lecture	
	16	Oxidation reduction titrations	Lecture	
	17	Experimental method, example	Lecture	

Head of the Department

Thevara

Date 2/11/2018

COURSE PLAN					
ACADEMIC YEAR 2016-17					
PROGRAMME	:	<i>B.Sc. Botany</i>	LECTURE HOURS	:	54
SEMESTER	:	4	CREDITS	:	3
SUBJECT TITLE	:	Advanced Bio-organic Chemistry	SUBJECT CODE	:	15U4CPCHE4. 2
COURSE TEACHERS	:	<i>Dr. Ramakrishnan S., Dr. Franklin John, Dr. Midhun Dominic</i>			
COURSE OUTCOMES (COs)	:	<p>CO1 - <i>To study the classification and properties of amino acids</i></p> <p>CO2 - <i>To study structure and functions of proteins</i></p> <p>CO3 - <i>To study classification, properties and structure of carbohydrates</i></p> <p>CO4 - <i>To study fundamentals of vitamins, hormones, steroids and lipids</i></p> <p>CO5 - <i>To study basics of natural products</i></p> <p>CO6 - <i>To study about heterocyclic compounds</i></p> <p>CO7 - <i>To study the fundamentals of chromatographic techniques</i></p>			

Dr. Ramakrishnan S				
Unit	No. of Sessions	Session Topic and Discussion Theme	Mode of teaching	COs
Amino acids and protein (12hrs)	1	Amino acids, Classification of amino acids	Lecture	CO1 CO2
	2	Chemical properties of amino acid	Lecture	
	3	Separation of amino acids	Lecture	
	4	Synthesis amino acids (glycine and alanine)	Lecture	
	5	Synthesis amino acids (phenyl alanine)	Lecture	
	6	Peptides, peptide bonds, polypeptides	Lecture	
	<i>First Internal examination</i>			
7	Proteins, Classification of proteins	Lecture		

	8	Prosthetic group, properties of proteins	Lecture	
	9	Denaturation of proteins	Lecture	
	10	Structure of proteins	Lecture	
	11	Structure of proteins	Lecture	
	12	Revision	Discussion	
Reference: <ol style="list-style-type: none"> 1. I. L. Finar, Organic Chemistry Vol 1 & 2, 6th edition, Pearson 2. K. S. Tewari, N. K. Vishnoi, A Text Book of Organic chemistry, 3rd edition, Vikas publishing House Pvt. Ltd ,2006. 3. Rastogi, Biochemistry, Tata McGraw –Hill Publication, 1996. 4. Dr. A.C. Deb, Fundamentals of Biochemistry. 				
Chromatographic Techniques (6 hrs)	13	Chromatographic Techniques, classifications	Lecture	CO7
	14	principle and uses: R _f values Column chromatography, Thin layer chromatography (TLC), Paper chromatography (PC):	Lecture	
	15	Gas chromatography(GC)	Lecture	
	<i>Second Internal examination</i>			
	16	High Performance Liquid chromatography (HPLC)	Lecture	
	17	Ion Exchange chromatography (IEC)	Lecture	
	18	Revision	Discussion	
Reference: <ol style="list-style-type: none"> 1. R. A. Day Junior, A.L. Underwood, Quantitative Analysis, 5th edn. Prentice Hall of India Pvt. Ltd. New Delhi, 1988(Chapters 17, 18). 1. Vogel's Text Book of Quantitative Chemical Analysis, J. Mendham, R. C. Denney, J.D. 2. Barnes, M. Thomas, 6th edn. Pearson Education (2003). 3. R. Gopalan, Analytical Chemistry, S. Chand and Co., New Delhi. 				
Dr. Franklin John				
Carbohydrates (10 hrs)	1	Carbohydrates, classification	Lecture	CO3
	2	preparation and properties of glucose	Lecture	
	3	preparation and properties of fructose	Lecture	
	4	preparation and properties of sucrose	Lecture	
	5	Haworth configuration of α -D glucose and β -D glucose, α -D fructose, β -D fructose, maltose and cellobiose	Lecture	
	6	Mutarotation	Lecture	
	<i>First Internal examination</i>			

	7	Conversion of glucose to fructose and vice-versa	Lecture		
	8	Structure of starch and cellulose	Lecture		
	9	Industrial applications of cellulose	Lecture		
	10	Revision	Discussion		
Vitamins, Steroids, Hormones and lipids (8hrs)	11	Vitamins, Structure and biological activity of Vitamin A, B and C	Lecture	CO4	
	12	Steroids	Lecture		
	13	Cholesterol	Lecture		
	14	Bile acids	Lecture		
	15	Lipids, simple and complex lipids	Lecture		
	Second Internal examination				
	16	Analysis of oils and fats – acid value, saponification value, iodine value	Lecture		
	17	Role MUFA and PUFA in preventing heart disease	Lecture		
	18	Revision	Discussion		
Reference					
<ol style="list-style-type: none"> 1. I. L. Finar, Organic Chemistry Vol 1 & 2, 6th edition, Pearson 2. K. S. Tewari, N. K. Vishnoi, A Text Book of Organic chemistry, 3rd edition, Vikas publishing House Pvt. Ltd ,2006. 3. Rastogi, Biochemistry, Tata McGraw –Hill Publication, 1996. 4. Dr. A.C. Deb, Fundamentals of Biochemistry. 					
Dr. Midhun Dominic					
Vitamins, Steroids, Hormones and lipids	1	Hormones	Lecture	CO4	
	2	Steroid hormones, Peptide hormones	Lecture		
	3	Amine hormones, Artificial hormones	Lecture		
	4	Revision	Discussion		
Natural Products (5hrs)	5	Terpinoids Essential oils: isolation	Lecture	CO5	
	6	isoprene rule Citral, geraniol, natural rubber	Lecture		
	Second Internal examination				
	7	Alkaloids: isolation, properties	Lecture		
	8	Structure of coniine, nicotine, piperine	Lecture		

	9	Revision	Discussion	
	References 1. I. L. Finar, Organic Chemistry Vol 1 & 2, 6th edition, Pearson 2. K. S. Tewari, N. K. Vishnoi, A Text Book of Organic chemistry, 3rd edition, Vikas publishing House Pvt. Ltd, 2006			
Heterocyclic Compounds (9 hrs)	10	Aromaticity	Lecture	CO6
	11	Huckel rule	Lecture	
	12	Structure of furan, indole, pyridine, pyrimidine, purine	Lecture	
	13	Aromaticity of furan, indole, pyridine	Lecture	
	14	Aromaticity of pyrimidine, purine	Lecture	
	15	Preparation, properties :furan, pyridine	Lecture	
	<i>Second Internal examination</i>			
	16	Preparation , properties : indole	Lecture	
	17	Preparation , properties : purine and pyrimidine	Lecture	
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Head of the Department

Thevara

Date

COURSE PLAN					
ACADEMIC YEAR 2016-17					
PROGRAMME	:	<i>B.Sc. Zoology</i>	LECTURE HOURS	:	54
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Dr. Ramakrishnan S				
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Amino acids and protein (12hrs)	1	Amino acids, Classification of amino acids	Lecture	CO1 CO2
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	3	Separation of amino acids	Lecture	
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	6	Peptides, peptide bonds, polypeptides	Lecture	
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Chromatographic Techniques (6 hrs)	13	Chromatographic Techniques, classifications	Lecture	CO7
	14	principle and uses: R _f values Column chromatography, Thin layer chromatography (TLC), Paper chromatography (PC):	Lecture	
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Dr. Franklin John				
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	13	Cholesterol	Lecture		
	14	Bile acids	Lecture		
	15	Lipids, simple and complex lipids	Lecture		
	Second Internal examination				
	16	Analysis of oils and fats – acid value, saponification value, iodine value	Lecture		
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Head of the Department

Thevara

Date